

Chemistry 338 2nd Semester Final Exam Review Sheet

Unit 1 – Matter & Measurement – Chapters 1 & 2

1. A student experimentally determines the density of an object to be 4.2 g/mL. If the accepted value is 5.0 g/mL, calculate the % error.

$$\% \text{ error} = \frac{\text{experimental} - \text{accepted}}{\text{accepted}} \times 100 \quad \% \text{ error} = \frac{4.2 - 5.0}{5.0} \times 100 = -16\% \text{ error}$$

Unit 2 – Periodic Table – Chapter 5

2.

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #	Charge
Tungsten W	$\frac{184}{74} W^{+2}$	74	74	110	72	184	+2

3. Using the following atomic radius values, predict the atomic radius of potassium (K):
Li – 145 pm, Na – 180 pm, K – ?, Rb – 235 pm
K – 220 pm

Unit 3 – Reactions/Compounds – Chapters 7 & 8

4. Classify the compound as ionic or covalent. Name the compound
a. Fe_2O_3 - iron (III) oxide (ionic – roman numerals for transition metals, charge ending to -ide)
b. Cl_2O_8 - dichlorine octaoxide (covalent – use prefixes, charge ending to -ide)

5. Classify the compound as ionic or covalent. Write the formula
a. Trisulfur Dinitride – S_3N_2 (covalent)
b. Zinc Hydroxide – $\text{Zn}(\text{OH})_2$ (ionic – look up ion charges and criss-cross) $\text{Zn}^{+2} \text{OH}^{-2}$
c. Chromium (III) nitrate – $\text{Cr}(\text{NO}_3)_3$ (ionic – criss-cross) $\text{Cr}^{+3} \text{NO}_3^{-1}$

6. Write the equation, balance it, and classify the type of reaction:
Iron (II) sulfate reacts with sodium hydroxide to yield iron(II) hydroxide and sodium sulfate
 $1 \text{FeSO}_4 + 2 \text{NaOH} \rightarrow 1 \text{Fe}(\text{OH})_2 + 1 \text{Na}_2\text{SO}_4$
double displacement (two ionic compounds reacting)

7. Balance the equation and classify the type of reaction:
 $1 \text{Cu} + 2 \text{AgNO}_3 \rightarrow 1 \text{Cu}(\text{NO}_3)_2 + 2 \text{Ag}$
Single displacement (element and a compound reacting)

Unit 4 – Moles – Chapter 3

8. Calculate the molar mass of $\text{Fe}_2(\text{SO}_4)_3$.
 $\text{Fe} (2)(55.9 \text{ g/mol}) = 111.8 \text{ g/mol}$
 $\text{S} (3)(32.1 \text{ g/mol}) = 96.3 \text{ g/mol}$
 $\text{O} (12)(16.0 \text{ g/mol}) = 192 \text{ g/mol}$
 $= 400.1 \text{ g/mol}$

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Michael Hoelscher

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1. Understanding the eBook 2014 Second Semester Chemistry Final
 - The Rise of Digital Reading 2014 Second Semester Chemistry Final
 - Advantages of eBooks Over Traditional Books
2. Identifying 2014 Second Semester Chemistry Final
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2014 Second Semester Chemistry Final
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2014 Second Semester Chemistry Final
 - Personalized Recommendations

- 2014 Second Semester Chemistry Final User Reviews and Ratings
- 2014 Second Semester Chemistry Final and Bestseller Lists

5. Accessing 2014 Second Semester Chemistry Final Free and Paid eBooks

- 2014 Second Semester Chemistry Final Public Domain eBooks
- 2014 Second Semester Chemistry Final eBook Subscription Services
- 2014 Second Semester Chemistry Final Budget-Friendly Options

6. Navigating 2014 Second Semester Chemistry Final eBook Formats

- ePUB, PDF, MOBI, and More
- 2014 Second Semester Chemistry Final Compatibility with Devices
- 2014 Second Semester Chemistry Final Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 2014 Second Semester Chemistry Final
- Highlighting and Note-Taking 2014 Second Semester Chemistry Final
- Interactive Elements 2014 Second Semester Chemistry Final

8. Staying Engaged with 2014 Second Semester Chemistry Final

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 2014 Second Semester Chemistry Final

9. Balancing eBooks and Physical Books 2014 Second Semester Chemistry Final

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 2014 Second Semester Chemistry Final

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 2014 Second Semester Chemistry Final

- Setting Reading Goals 2014 Second Semester Chemistry Final
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 2014 Second Semester Chemistry Final

- Fact-Checking eBook Content of 2014 Second Semester Chemistry Final

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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