

2. (a) An aqueous solution of a salt was added to excess sodium trioxocarbonate(IV) solution in a test tube. There was effervescence, the test tube became warm and a white precipitate was observed. State three inferences that could be drawn from these observations.

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- (b) Consider the following salts:

$\text{NH}_4\text{Cl}$ ;  $\text{PbSO}_4$ ;  $\text{NaHCO}_3$ ; and  $\text{Cu}(\text{NO}_3)_2$

Select from the list, the salt(s) which:

- (i) would not readily dissolve in water;
- 
- (ii) produce(s) effervescence with dilute mineral acids;
- 
- (iii) decompose(s) on heating;
- 
- (iv) dissolve(s) in water to form an alkaline solution;
- 
- (v) sublime(s) on heating.
-

# 2014 Waec Chemistry Practicals Questions And Answers For Alternative B

**L Towne**



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