

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the  $K_a$  and  $K_b$  of acids and bases and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                   |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. $K_a$                    | 6. $K_b$          |
| 7. ionization     | 8. $K_a$                    | 9. $K_b$          |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the numerical value(s) of the  $K_a$  and  $K_b$  for the acids and bases listed in the table. Also indicate whether the acid or base is strong or weak.

1.  $\text{HCl}$  is a strong acid. Write the equation and/or the numerical value(s) of the  $K_a$  and  $K_b$  for the acid and base.
2.  $\text{HCl}$  is a strong acid. Write the equation and/or the numerical value(s) of the  $K_a$  and  $K_b$  for the acid and base.
3.  $\text{HCl}$  is a strong acid. Write the equation and/or the numerical value(s) of the  $K_a$  and  $K_b$  for the acid and base.
4.  $\text{HCl}$  is a strong acid. Write the equation and/or the numerical value(s) of the  $K_a$  and  $K_b$  for the acid and base.
5.  $\text{HCl}$  is a strong acid. Write the equation and/or the numerical value(s) of the  $K_a$  and  $K_b$  for the acid and base.
6.  $\text{HCl}$  is a strong acid. Write the equation and/or the numerical value(s) of the  $K_a$  and  $K_b$  for the acid and base.
7.  $\text{HCl}$  is a strong acid. Write the equation and/or the numerical value(s) of the  $K_a$  and  $K_b$  for the acid and base.
8.  $\text{HCl}$  is a strong acid. Write the equation and/or the numerical value(s) of the  $K_a$  and  $K_b$  for the acid and base.
9.  $\text{HCl}$  is a strong acid. Write the equation and/or the numerical value(s) of the  $K_a$  and  $K_b$  for the acid and base.
10.  $\text{HCl}$  is a strong acid. Write the equation and/or the numerical value(s) of the  $K_a$  and  $K_b$  for the acid and base.

|     |              |
|-----|--------------|
| 1.  | $\text{HCl}$ |
| 2.  | $\text{HCl}$ |
| 3.  | $\text{HCl}$ |
| 4.  | $\text{HCl}$ |
| 5.  | $\text{HCl}$ |
| 6.  | $\text{HCl}$ |
| 7.  | $\text{HCl}$ |
| 8.  | $\text{HCl}$ |
| 9.  | $\text{HCl}$ |
| 10. | $\text{HCl}$ |

## 19 2 Hydrogen Ions And Acidity Answer Key

**SJ Ball**



## 19 2 Hydrogen Ions And Acidity Answer Key:

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