

Experiment 3

Advance Study Assignment: Resolution of Matter into Pure Substances. II. Fractional Crystallization

1. Using Figure 3.1, determine

a. the number of grams of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ that will dissolve in 100 g of H_2O at 100°C.

$$(63.55) + (25.06) + (100.0) = 188.6 \text{ g of CuSO}_4$$

$$180.0 \text{ g CuSO}_4 \cdot 5\text{H}_2\text{O}$$

b. the number of grams of water required to dissolve 4.0 g of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ at 100°C. (Hint: your answer to Problem 1(a) gives you the needed conversion factor for g $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ to g H_2O .)

$$4 \text{ g CuSO}_4 \cdot 5\text{H}_2\text{O} \times \frac{100.0 \text{ g H}_2\text{O}}{180.0 \text{ g CuSO}_4 \cdot 5\text{H}_2\text{O}} = 2.222 \text{ g H}_2\text{O}$$

c. the number of grams of water required to dissolve 18 g KNO_3 at 100°C.

$$18 \text{ g KNO}_3 \times \frac{100.0 \text{ g H}_2\text{O}}{25.0 \text{ g KNO}_3} = 72.0 \text{ g H}_2\text{O}$$

$$72.0 \text{ g H}_2\text{O}$$

d. the number of grams of water required to dissolve a mixture containing 18 g KNO_3 and 4.0 g $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$, assuming that the solubility of one substance is not affected by the presence of another.

$$18 \text{ g KNO}_3 \times \frac{100.0 \text{ g H}_2\text{O}}{25.0 \text{ g KNO}_3} = 72.0 \text{ g H}_2\text{O}$$

(continued on following page)

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**Susan A. Weiner, Peters, Edward I.
Peters**

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