

Stoichiometry & the mole notes

the mole concept

the mole (mol) is the SI unit for measuring the amount of a substance.

Avogadro's number:

$$1 \text{ mol} = 6.022 \times 10^{23} \text{ particles}$$

(particles can be atoms, molecules, or formula units)

molar mass:

the mass of one mole of a substance in grams.

found by adding the atomic masses from the periodic table

example:

$$\text{H}_2\text{O} (2 \cdot 1.008 \text{ g/mol}) + (1 \cdot 16.00 \text{ g/mol}) = 18.016 \text{ g/mol}$$

Example Problems

1. how many atoms are in 3.5 moles of Fe?

$$3.5 \text{ mol Fe} \cdot \frac{6.022 \times 10^{23} \text{ atoms}}{1 \text{ mol}} = 2.11 \times 10^{24} \text{ atoms}$$

2. what is the mass of 2.00 moles of NaCl?

Find molar mass:

$$\text{Na} : 22.99 \text{ g/mol}$$

$$\text{Cl} : 35.45 \text{ g/mol}$$

$$\text{Total} : 58.44 \text{ g/mol}$$

Convert:

$$2.00 \text{ mol} \cdot \frac{58.44 \text{ g}}{1 \text{ mol}} = 116.88 \text{ g}$$



Stoichiometry & The Mole

Conversions Involving the mole

conversion

particles \leftrightarrow moles

Factor to use:

$$1 \text{ mol} = 6.022 \times 10^{23} \text{ particles}$$

moles \leftrightarrow mass

$$1 \text{ mol} = \text{molar mass (g)}$$

moles \leftrightarrow volume (L) (at STP)

$$1 \text{ mol} = 22.4 \text{ L of gas}$$

Stoichiometry: the study of quantitative relationships in chemical reactions

- Balanced chemical Equations Show the mole ratio of Reactants & products
- mole Ratios Come from coefficients in a balanced equation.

Stoichiometry Steps:

- Balance the equation
- convert given quantity into moles (if necessary)
- Use Mole Ratio from balanced equation
- Convert moles back to desired units (grams, liters, etc.)

Example:



question: how many grams of H_2O are produced from 5.00 moles of O_2 ?

Step 1: Use Mole Ratio:

$$5.00 \text{ mol O}_2 \cdot \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} = 10.0 \text{ mol H}_2\text{O}$$

Step 2: Convert to grams

$$10.0 \text{ mol H}_2\text{O} \cdot \frac{18.0 \text{ g}}{1 \text{ mol}} = 180. \text{ Kg H}_2\text{O}$$

Limiting Reactant & Percent Yield

Limiting Reactant (LR): the reactant that runs out first, limiting product formation

Excess Reactant: the reactant that remains after the reaction is complete

Steps to identify LR

- Convert Reactants to moles

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- use mole ratios to do

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Rosina Ehmann

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