



**821F  
921F  
Tier 4  
Wheel Loader**

**OPERATOR'S MANUAL**

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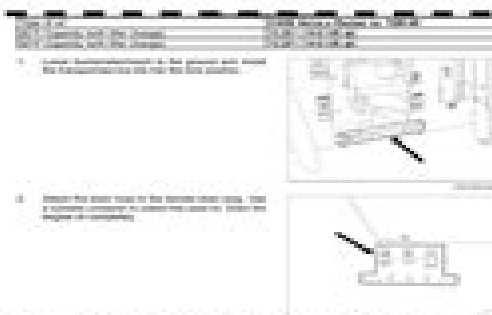
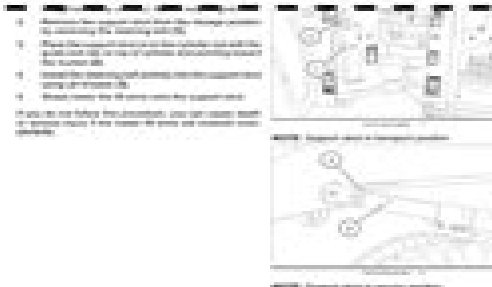
**CASE**  
CONSTRUCTION



**SAFETY PRECAUTIONS**  
The operator must read and understand the safety precautions in this manual before operating the machine. The safety precautions in this manual are intended to provide the operator with the information needed to operate the machine safely and to prevent injury or death. The operator must read and understand the safety precautions in this manual before operating the machine.

**Safety precautions specific to attachment and coupling**

- Always inspect the attachment and coupling for damage before use. If the attachment or coupling is damaged, do not use it.
- Use the correct attachment and coupling for the machine. Do not use an attachment or coupling that is not approved by the manufacturer.
- Do not attach or detach an attachment or coupling while the machine is in motion.
- Do not attach or detach an attachment or coupling while the machine is on a slope.
- Do not attach or detach an attachment or coupling while the machine is on a wet or icy surface.
- Do not attach or detach an attachment or coupling while the machine is on a soft or uneven surface.
- Do not attach or detach an attachment or coupling while the machine is on a hard or uneven surface.
- Do not attach or detach an attachment or coupling while the machine is on a hard or uneven surface.
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# Case 821f 921f Wheel Loader Operators Manual

**Liying Dong**



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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is  $2\text{Cl}^- + 2\text{Ag}^+ \rightarrow 2\text{AgCl}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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$1\text{K}_3\text{PO}_4 + 3\text{Al}(\text{OH})_3 + 3\text{HNO}_3 \rightarrow 3\text{KNO}_3 + 3\text{H}_3\text{PO}_4 + 3\text{H}_2\text{O}$  net ionic  $\text{Al}^{3+} + 3\text{OH}^- \rightarrow \text{Al}(\text{OH})_3$   $\text{Al}^{3+} + 3\text{OH}^- \rightarrow \text{Al}(\text{OH})_3$   $\text{Al}^{3+} + 3\text{OH}^- \rightarrow \text{Al}(\text{OH})_3$

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oct 2 2019 the net ionic equation is a chemical equation for a reaction that lists only those species participating in the reaction the net ionic equation is commonly used in acid base

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reactant side copper on the product side are neutral

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds  $\text{Cl}^-$  aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver  $\text{I}^+$  ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4  $13 \text{Ag}^+ \text{aq} + 2\text{F}^- \text{aq} + 2\text{NH}_4^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s} + 2\text{NH}_4^+ \text{aq} + 2\text{F}^- \text{aq}$  because two  $\text{NH}_4^+$  aq

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