

4**ATOMIC STRUCTURE****Chapter Test A****A. Matching**

Match each description in *Column B* with the correct term in *Column A*. Write the letter of the correct description on the line.

Column A

J 1. proton
G 2. atom
A 3. mass number
C 4. atomic mass unit
F 5. electron
E 6. isotopes
D 7. atomic number
B 8. atomic mass
H 9. period
I 10. neutron

Column B

a. the total number of protons and neutrons in the nucleus of an atom
 b. the weighted average mass of the atoms in a naturally occurring sample of an element
 c. $\frac{1}{12}$ the mass of a carbon-12 atom
 d. the number of protons in the nucleus of an element
 e. atoms with the same number of protons but different numbers of neutrons
 f. negatively charged subatomic particle
 g. the smallest particle of an element that retains its identity in a chemical reaction
 h. a horizontal row of the periodic table
 i. subatomic particle with no charge
 j. positively charged subatomic particle

B. Multiple Choice

Choose the best answer and write its letter on the line.

11. Which of the following is *not* a part of Dalton's atomic theory?
 a. All elements are composed of atoms.
 b. Atoms of the same element are alike.
 c. Atoms are always in motion.
 d. Atoms that combine do so in simple whole-number ratios.

12. The nucleus of an atom is.
 a. negatively charged and has a low density.
 b. negatively charged and has a high density.
 c. positively charged and has a low density.
 d. positively charged and has a high density.

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Albert Josiah Harris,Edward R. Sipay

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