

## Chemistry Final Exam Review Questions

### Useful Values and Equations

Some of the following values and equations may be useful for answering some of the questions and problems.

$$Q_F = 1.84 Q_C + 12$$

speed of light =  $3.00 \times 10^8$  m/s

2000

$R = 0.082057 \text{ L-atm / deg-mole}$

Planck's constant =  $6.63 \times 10^{-34} \text{ J.s}$

15. *Journal of the American Medical Association*, 2000; 283: 2689-2694.

### Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- One of the following does **not** describe gases. Which one is it?
  - Expands infinitely.
  - Rigid shape.
  - Easily compressed.
  - Relatively low densities.
- The following properties describe the element, mercury (Hg). Which one is a **chemical** property?
  - Its density is  $5.43 \text{ g/cm}^3$ .
  - It is a silvery-white metal.
  - It is a liquid at 298 K.
  - Mercury(II) sulfide forms when it is exposed to sulfide ions.
- All of the following are properties of  $\text{SO}_2$ . Which one is a **physical** property?
  - It reacts with oxygen to form  $\text{SO}_3^{2-}$ .
  - It condenses to a colorless liquid at  $14^\circ\text{F}$  and atmospheric pressure.
  - It is a strong reducing agent.
  - It reacts with water and oxygen to form an acid.
- Which of the following numbers has 4 significant figures?
  - 0.04309
  - 0.0430
  - 0.0431
  - 0.43980
- How many significant figures should there be in the answer to the following math equation?
$$2.834 + 5.71520 + 2.12 + 178.1 + 250.2619$$
  - 3
  - 4
  - 5
  - 6
- How many significant figures should there be in the answer to the following math equation?
$$23.413 \text{ g} + (2.15 \text{ cm} \times 1.1 \text{ cm} \times 3.73 \text{ cm})$$
  - 2
  - 3
  - 4
  - 5
- At  $25^\circ\text{C}$ , one  $\text{cm}^3$  of iron has a mass of 7.62 grams. What volume of iron is occupied by 125 grams of iron?
  - $0.82 \text{ cm}^3$
  - $8.2 \text{ cm}^3$
  - $0.164 \text{ cm}^3$
  - $16.4 \text{ cm}^3$
- A cube of lead has a volume of  $6.61 \text{ cm}^3$ . Its mass is 75.0 grams. What is its density?
  - $3.85 \text{ g/cm}^3$
  - $7.41 \text{ g/cm}^3$
  - $11.3 \text{ g/cm}^3$
  - $22.6 \text{ g/cm}^3$

# 2014 Spring Final Chemistry Review Answers

**EW Minium**



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Research in Chemistry Education Liliana Mammino, Jan Apotheker, 2021-05-17 This volume emphasizes the role of chemical education for development and in particular for sustainable development in Africa by sharing experiences among specialists across the African continent and with specialists from other continents It considers all areas and levels of chemistry education gives specific attention to known major challenges and encourages explorations of novel approaches The chapters in this book describe new teaching approaches approach explorations and in class activities analyse educational challenges and possible ways of addressing them and explore cross discipline possibilities and their potential benefits for chemistry education This makes the volume an up to date compendium for chemistry educators and educational researchers worldwide

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*Scientific and Technical Aerospace Reports* ,1970 Lists citations with abstracts for aerospace related reports obtained from world wide sources and announces documents that have recently been entered into the NASA Scientific and Technical Information Database

*Journal of Developmental Education* ,2017

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