

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and calculate  $pK_a$  and  $pK_b$  for weak acids and bases.
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and calculate  $pK_a$  and  $pK_b$  for weak acids and bases.
3. Calculate the  $pH$  of weak acid and base solutions.

## Vocabulary

- |                 |                                      |                 |
|-----------------|--------------------------------------|-----------------|
| acid            | acid dissociation constant ( $K_a$ ) | acid ionization |
| acidic solution | acidic solution                      | acidic solution |
| acidic solution | acidic solution                      | acidic solution |

## Key Equations

- |   |  |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

## Part A Completion

Write the chemical formula for each of the following acids, bases, and salts, showing the correct number of atoms of each element in the formula.

1. Acetic acid,  $CH_3COOH$
2. Hydrofluoric acid,  $HF$
3. Hydrochloric acid,  $HCl$
4. Hydrobromic acid,  $HBr$
5. Hydroiodic acid,  $HI$
6. Nitric acid,  $HNO_3$
7. Sulfuric acid,  $H_2SO_4$
8. Phosphoric acid,  $H_3PO_4$
9. Acetic acid,  $CH_3COOH$
10. Hydrofluoric acid,  $HF$
11. Hydrochloric acid,  $HCl$
12. Hydrobromic acid,  $HBr$
13. Hydroiodic acid,  $HI$
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16. Phosphoric acid,  $H_3PO_4$

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**JR Anderson**



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### **Table of Contents 19 2 Hydrogen Ions And Acidity Pearson**

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Pearson
  - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Pearson
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Pearson
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 2 Hydrogen Ions And Acidity Pearson
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Pearson

- Personalized Recommendations
- 19 2 Hydrogen Ions And Acidity Pearson User Reviews and Ratings
- 19 2 Hydrogen Ions And Acidity Pearson and Bestseller Lists
- 5. Accessing 19 2 Hydrogen Ions And Acidity Pearson Free and Paid eBooks
  - 19 2 Hydrogen Ions And Acidity Pearson Public Domain eBooks
  - 19 2 Hydrogen Ions And Acidity Pearson eBook Subscription Services
  - 19 2 Hydrogen Ions And Acidity Pearson Budget-Friendly Options
- 6. Navigating 19 2 Hydrogen Ions And Acidity Pearson eBook Formats
  - ePub, PDF, MOBI, and More
  - 19 2 Hydrogen Ions And Acidity Pearson Compatibility with Devices
  - 19 2 Hydrogen Ions And Acidity Pearson Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Pearson
  - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Pearson
  - Interactive Elements 19 2 Hydrogen Ions And Acidity Pearson
- 8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Pearson
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Pearson
- 9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Pearson
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Pearson
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Pearson
  - Setting Reading Goals 19 2 Hydrogen Ions And Acidity Pearson
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Pearson

- Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Pearson
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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