

**Class : 11**

Register  
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## SECOND REVISION EXAMINATION - 2025

Time Allowed : 3.00 Hours]

### CHEMISTRY PART - I

[Max. Marks : 70

15x1=15

I. Answer the following:

- 7.5 g of a gas occupies a volume of 5.6 litres at 0°C and 1 atm pressure. The gas is  
a) NO                                      b)  $N_2O$                                       c) CO                                      d)  $CO_2$
- Which of the following pairs of d-orbitals will have electron density along the axes?  
a)  $d_{xz}, d_{yz}$                                       b)  $d_{xy}, d_{yz}$                                       c)  $d_{xz}, d_{xy}$                                       d)  $d_{xy}, d_{yz}$
- The element with positive electron gain enthalpy is  
a) Hydrogen                                      b) Sodium                                      c) Argon                                      d) Fluorine
- The hardness of water can be determined by volumetrically using the reagent  
a) Sodium thio sulphate                                      b) Potassium Permanganate  
c) Hydrogen Peroxide                                      d) EDTA
- \_\_\_\_\_ is used in Cancer therapy.  
a)  $^{88}Sr$                                       b)  $^{90}Sr$                                       c)  $^{86}Sr$                                       d)  $^{11}Be$
- Maximum deviation from ideal gas is expected from  
a)  $CH_4(g)$                                       b)  $NH_3(g)$                                       c)  $H_2(g)$                                       d)  $N_2(g)$
- The values of  $\Delta H$  and  $\Delta S$  for a reaction are respectively 30  $KJ\cdot mol^{-1}$  and 100  $JK^{-1}\cdot mol^{-1}$ . Then the temperature above which the reaction will become spontaneous is  
a) 300 K                                      b) 30 K                                      c) 100 K                                      d) 20°C
- In a chemical equilibrium, the rate constant for the forward reaction is  $2.5 \times 10^3$  and the equilibrium constant is 50. The rate constant for the reverse reaction is  
a) 11                                      b) 5                                      c)  $2 \times 10^3$                                       d)  $2 \times 10^4$
- Assertion :** An ideal solution obeys Raoult's law.  
**Reason :** In an ideal solution, solvent-solvent as well as solute - solute interactions are similar to solute - solvent interactions.  
a) Both Assertion and reason are true and reason is the correct explanation of assertion.  
b) Both Assertion and reason are true but reason is not the correct explanation of assertion.  
c) Assertion is true but reason is false                                      d) Both assertion and reason are false.

# 2mid Year Question Papers

**Jin-Ying Zhang**



## 2mid Year Question Papers:

CBSE Class 10 Maths Standard Chapterwise Previous Year Question Bank (PYQ) For 2026 Exam Nodia,2025-12-27 CBSE Mathematics Chapter Wise Question Bank covers all questions from 63 sets of CBSE Maths previous year papers 2006 2025 Based on the latest CBSE syllabus and exam pattern It includes MCQs very short short long and competency based case study questions in a chapter wise format This book focuses on concept clarity problem solving skills **Oswaal CBSE**

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Education Today ,1990 **Paper Trade Journal** ,1961 Accounts and Papers of the House of Commons Great Britain. Parliament. House of Commons,1872 **Leadership for Professional Learning** Sue Swaffield,Philip E. Poekert,2022-12-26 This book brings together a collection of inquiries into the connections between educational leadership understood as an activity that can be performed by both educators and students and professional learning understood as an activity undertaken by educators to improve teaching and learning within educational settings The book is framed by two reviews of the academic literature which together provide a broad overview of the published literature as well as a more targeted look at where this work intersects with issues of educational equity The remaining chapters which include both conceptual and empirical pieces explore leadership for professional learning from multiple vantage points including student leadership teacher leadership senior leadership and shared leadership across roles Collectively the chapters contribute to challenging the commonly accepted notion that the exercise of leadership is the sole purview of those in positions of status and honoring the complexity of interactions among students teachers and senior leaders that influence teaching and learning outcomes In so doing they inform both future practice and research All but one of the chapters in this book were originally published in the journal Professional Development in Education Southern Pulp and Paper ,1984 **Parliamentary**

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