

SECTION B

3. Answer ALL parts.

a. For **BOTH** $\text{Na}[\text{CoCl}_4]$ and $\text{Na}_3[\text{CoCl}_6]$:

- i) Calculate the Crystal Field Stabilization Energy. (3 x 2)
- ii) Determine if Jahn-Teller distortions will be observed. (2 x 2)
- iii) Derive the hybridization at the metal center. (2 x 2)

b. i) Given that Cl^- is a π -donor ligand, draw the molecular orbital (MO) diagram for $\text{Na}_3[\text{CoCl}_6]$. Label the metal orbitals, the ligand group orbitals and the molecular orbitals with the appropriate symmetry labels. Indicate which molecular orbitals are bonding, antibonding and non-bonding. (9)

ii) Fill electrons into the orbitals and label the transition that corresponds to Δ . (2)

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