

2. Tertiary acids -- just like binary acids, tertiary acids must begin with hydrogen. These acids also contain a polyatomic ion.

Polyatomic ion	Name of ion	Polyatomic ion	Name of ion
OH^-	Hydroxide	PO_4^{3-}	Phosphate
O_2^{2-}	Peroxide	ClO_4^-	perchlorate
$\text{C}_2\text{O}_4^{2-}$	Acetate	ClO_3^-	chlorate
S_2^{2-}	Sulfide	ClO_2^-	chlorite
S_2^{2-}	sulfate	ClO^-	hypochlorite

Tertiary acids are three element acids. Most tertiary acids are OXYACIDS containing hydrogen, oxygen and other element.

Naming:

- Name the anion within the acid
- Change the anion ending to one of the following:

Anion ending	In acid, replace with:
-ide	-ous acid
-ate	-ic acid

Formula	anion	Tertiary compound name	Tertiary acid name
H_2SO_3	SO_3^{2-}	Hydrogen Sulfite	Sulfurous acid
H_2SO_4	SO_4^{2-}	Hydrogen sulfate	Sulfuric acid
HNO_2	NO_2^-	Hydrogen Nitrite	Nitrous acid
HNO_3	NO_3^-	Hydrogen Nitrate	Nitric acid
HClO_4	ClO_4^-	perchlorate	Perchloric acid
HClO_3	ClO_3^-	chlorate	Chloric acid
HClO_2	ClO_2^-	chlorite	Chlorous acid
HClO	ClO^-	hypochlorite	Hypochlorous acid

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Charles H. Corwin



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