

Chemistry

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(Chapter 1)(Solutions) (Exercises)
(Class 12)

Question 1.26:

If the density of some lake water is 1.25 g mL^{-1} and contains 92 g of Na^+ ions per kg of water, calculate the molality of Na^+ ions in the lake.

Answer 1.26:

Number of moles present in 92 g of Na^+ ions = $\frac{92 \text{ g}}{23 \text{ g mol}^{-1}} = 4 \text{ mol}$

Therefore, molality of Na^+ ions in the lake = $\frac{4 \text{ mol}}{1 \text{ kg}} = 4 \text{ m}$

Question 1.27:

If the solubility product of CuS is 6×10^{-16} , calculate the maximum molarity of CuS in aqueous solution.

Answer 1.27:

Solubility product of CuS , $K_{sp} = 6 \times 10^{-16}$ Let s be the solubility of CuS in mol L^{-1} .



$$K_{sp} = [\text{Cu}^{2+}][\text{S}^{2-}]$$

Now,

$$= s \times s = s^2$$

Then, we have, $K_{sp} = s^2 = 6 \times 10^{-16}$

$$\Rightarrow s = \sqrt{6 \times 10^{-16}} = 2.45 \times 10^{-8} \text{ mol L}^{-1}$$

Hence, the maximum molarity of CuS in an aqueous solution is $2.45 \times 10^{-8} \text{ mol L}^{-1}$.

Question 1.28:

Calculate the mass percentage of aspirin ($\text{C}_9\text{H}_8\text{O}_4$) in acetonitrile (CH_3CN) when 6.5 g of $\text{C}_9\text{H}_8\text{O}_4$ is dissolved in 450 g of CH_3CN .

Answer 1.28:

6.5 g of $\text{C}_9\text{H}_8\text{O}_4$ is dissolved in 450 g of CH_3CN .

Then, total mass of the solution = $(6.5 + 450) \text{ g} = 456.5 \text{ g}$

Therefore, mass percentage of $\text{C}_9\text{H}_8\text{O}_4 = \frac{6.5}{456.5} \times 100 \% = 1.424\%$

Question 1.29:

Nalorphene ($\text{C}_{19}\text{H}_{21}\text{NO}_3$), similar to morphine, is used to combat withdrawal symptoms in narcotic users. Dose of nalorphene generally given is 1.5 mg. Calculate the mass of $1.5 \times 10^{-3} \text{ m}$ aqueous solution required for the above dose.

Answer 1.29:

The molar mass of Nalorphene ($\text{C}_{19}\text{H}_{21}\text{NO}_3$) is given as:

$$19 \times 12 + 21 \times 1 + 1 \times 14 + 3 \times 16 = 311 \text{ g mol}^{-1}$$

In $1.5 \times 10^{-3} \text{ m}$ aqueous solution of nalorphene, 1 kg (1000 g) of water contains $1.5 \times 10^{-3} \text{ mol}$
 $= 1.5 \times 10^{-3} \times 311 \text{ g} = 0.4665 \text{ g}$

Therefore, total mass of the solution = $(1000 + 0.4665) \text{ g} = 1000.4665 \text{ g}$

This implies that the mass of the solution containing 0.4665 g of nalorphene is 1000.4665 g.

Therefore, mass of the solution containing 1.5 mg of nalorphene is: $\frac{1000.4665 \times 1.5 \times 10^{-3}}{0.4665} \text{ g} = 3.22 \text{ g}$

Hence, the mass of aqueous solution required is 3.22 g.

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web feb 3 2006 4 reactor ap1000 design control document tier 2 material 4 4 3 revision 19 have shown that fuel melting will not occur for achievable local burnups up to 75 000 mwd mtu reference 81 the nrc has approved design evaluations up to 60 000 mwd mtu in reference 81 and up to 62 000 mwd mtu in references 9 and 88

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web 1 introduction and general description of plant design control document criteria referenced ap1000 section criteria position clarification summary description of exceptions eps gw gl 700 1a 2 revision 1 d 1 conforms guidance applies only to the class 1e dc and ups system since the ap1000 ac power system is a nonsafety related

16 technical specifications ap1000 design control document - Feb 27 2022

web 16 technical specifications ap1000 design control document tier 2 material 16 1 2 revision 19 this set of technical specifications is intended to be used as a guide in the development of the plant specific technical specifications the preliminary information originally provided in

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web reactor ap1000 design control document tier 2 material 4 3 1 revision 16 4 3 nuclear design 4 3 1 design basis this section describes the design bases and functional requirements used in the nuclear design of the fuel and reactivity control system and relates these design bases to the general design criteria gdc

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web the design engineering and other information contained in this document have been prepared by or on behalf of the westinghouse electric company llc in connection with its application to the united states nuclear regulatory commission nrc for design certification of the ap1000 passive nuclear plant design pursuant to title 10 code of

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web jan 5 2023 project overview on march 28 2002 the nrc staff initiated its formal review of the ap1000 pressurized water reactor pwr design when westinghouse electric company submitted its application for final design approval fda and standard

design certification for the ap1000 design

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power flow operating map boiling water

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revision 19 the analysis method is not recommended for complex equipment that cannot be modeled to adequately predict its
response 3 10 2 1 3 combination of test and analysis

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web structures systems and components in the ap1000 are classified according to nuclear safety classification quality groups
seismic category and codes and standards this section provides the methodology used for safety related and seismic
classification of ap1000 structures systems and components

7 instrumentation and controls ap1000 design control document - Jan 09 2023

web 7 instrumentation and controls ap1000 design control document tier 2 material 7 7 3 revision 18 7 7 1 1 reactor power
control system automatic reactor power and power distribution control are the basic functions of the reactor power control
system they are achieved by varying the position of the control rods separate

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web this design control document dcd for the westinghouse ap1000 simplified passive advanced light water reactor plant is
incorporated by reference into the design certification rule for the ap1000 design section ii a of appendix d to 10 cfr part 52

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11 2 revision 1 and maintenance schedules for that equipment the maximum and minimum conditions identified as the
abnormal condition are based on the design limits for the affected areas

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web equipment and systems ap1000 design control document nrc staff approval is required prior to implementing a change
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tier 1 information intro 1

