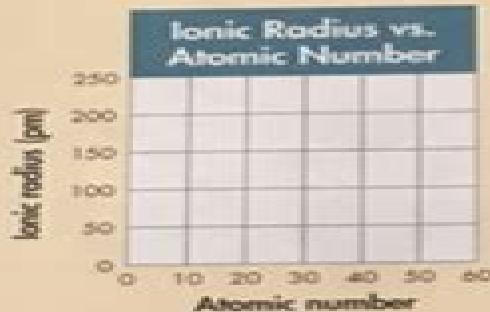


## Quick Lab

Purpose: To use a graph to identify period and group trends.

### Materials

- graph paper
- pencil



## Periodic Trends in Ionic Radii

### Procedure

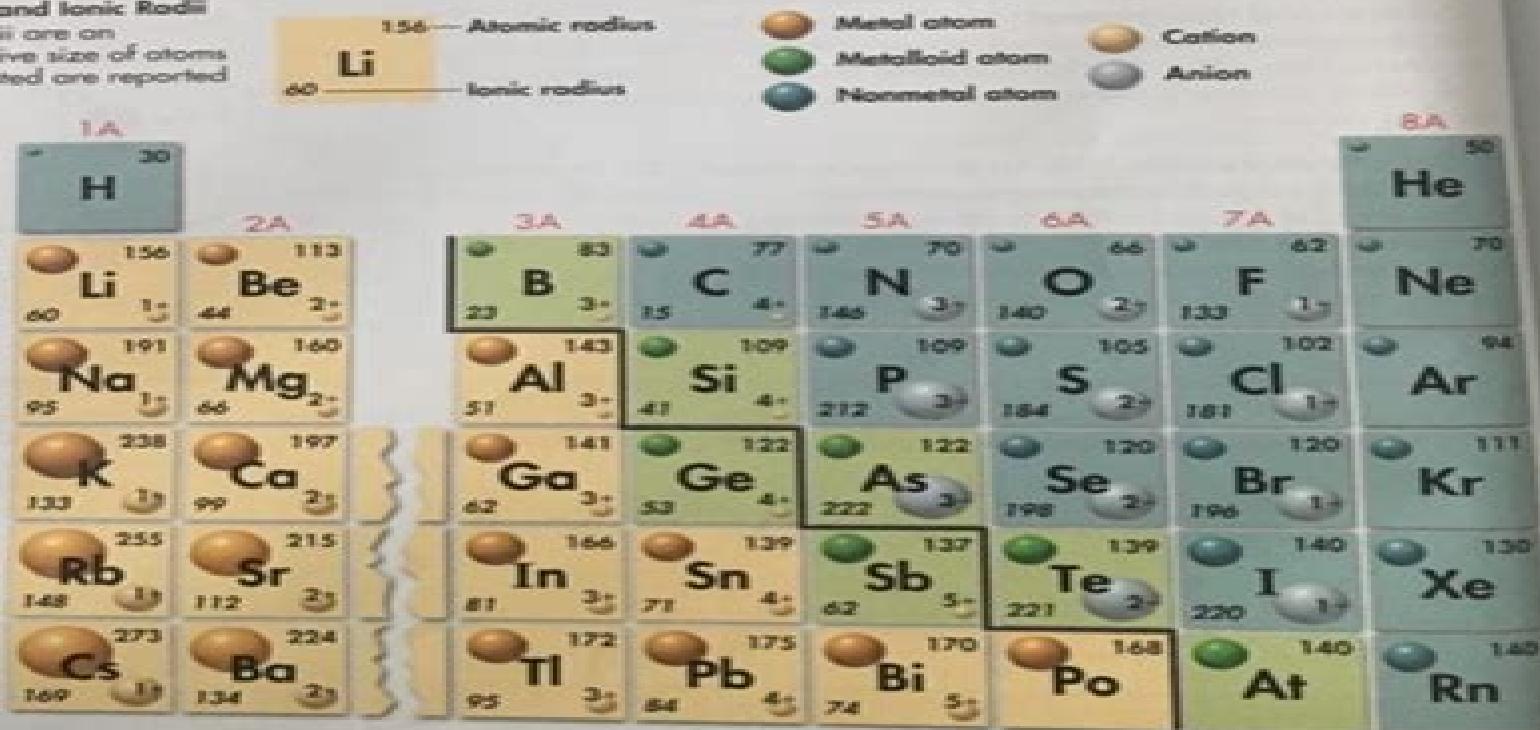
Use the data presented in Figure 6.23 to plot ionic radius versus atomic number.

### Analyze and Conclude

- Compare:** How does the size change when an atom forms a cation and when an atom forms an anion?
- Describe:** How do the ionic radii vary within a group of metals? How do they vary within a group of nonmetals?
- Describe:** What is the shape of a portion of the graph that corresponds to one period?
- Compare and Contrast:** Is the trend across a period similar or different for Periods 2, 3, 4, and 5?
- Explain:** Propose explanations for the trends you have described for ionic radii within groups and across periods.

Figure 6.23: Atomic and Ionic Radii

Atomic and ionic radii are an indication of the relative size of atoms and ions. The data listed are reported in picometers (pm).



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**JG Myers**

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## **Table of Contents Answers Quick Lab Periodic Trends In Ionic Radii**

1. Understanding the eBook Answers Quick Lab Periodic Trends In Ionic Radii
  - The Rise of Digital Reading Answers Quick Lab Periodic Trends In Ionic Radii
  - Advantages of eBooks Over Traditional Books
2. Identifying Answers Quick Lab Periodic Trends In Ionic Radii
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answers Quick Lab Periodic Trends In Ionic Radii
  - User-Friendly Interface

4. Exploring eBook Recommendations from Answers Quick Lab Periodic Trends In Ionic Radii
  - Personalized Recommendations
  - Answers Quick Lab Periodic Trends In Ionic Radii User Reviews and Ratings
  - Answers Quick Lab Periodic Trends In Ionic Radii and Bestseller Lists
5. Accessing Answers Quick Lab Periodic Trends In Ionic Radii Free and Paid eBooks
  - Answers Quick Lab Periodic Trends In Ionic Radii Public Domain eBooks
  - Answers Quick Lab Periodic Trends In Ionic Radii eBook Subscription Services
  - Answers Quick Lab Periodic Trends In Ionic Radii Budget-Friendly Options
6. Navigating Answers Quick Lab Periodic Trends In Ionic Radii eBook Formats
  - ePUB, PDF, MOBI, and More
  - Answers Quick Lab Periodic Trends In Ionic Radii Compatibility with Devices
  - Answers Quick Lab Periodic Trends In Ionic Radii Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answers Quick Lab Periodic Trends In Ionic Radii
  - Highlighting and Note-Taking Answers Quick Lab Periodic Trends In Ionic Radii
  - Interactive Elements Answers Quick Lab Periodic Trends In Ionic Radii
8. Staying Engaged with Answers Quick Lab Periodic Trends In Ionic Radii
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answers Quick Lab Periodic Trends In Ionic Radii
9. Balancing eBooks and Physical Books Answers Quick Lab Periodic Trends In Ionic Radii
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answers Quick Lab Periodic Trends In Ionic Radii
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Answers Quick Lab Periodic Trends In Ionic Radii
  - Setting Reading Goals Answers Quick Lab Periodic Trends In Ionic Radii
  - Carving Out Dedicated Reading Time

---

- 12. Sourcing Reliable Information of Answers Quick Lab Periodic Trends In Ionic Radii
  - Fact-Checking eBook Content of Answers Quick Lab Periodic Trends In Ionic Radii
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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