

Assignment #3 – Compound Names and Formulas Single and Multi-valent ions

A. Write the correct formula for the following compounds.

1. gallium oxide	<u>Ga₂O₃</u>	11. silicon oxide	<u>SiO₂</u>
2. mercury (I) chloride	<u>HgCl</u>	12. aluminum bromide	<u>AlBr₃</u>
3. tin (IV) bromide	<u>SnBr₄</u>	13. tin (II) sulphide	<u>SnS</u>
4. indium oxide	<u>In₂O₃</u>	14. germanium bromide	<u>GeBr₄</u>
5. cadmium sulphide	<u>CdS</u>	15. nickel (II) oxide	<u>NiO</u>
6. magnesium oxide	<u>MgO</u>	16. strontium astatide	<u>SrAt₂</u>
7. potassium nitride	<u>K₃N</u>	17. aluminum oxide	<u>Al₂O₃</u>
8. iron (III) sulphide	<u>Fe₂S₃</u>	18. copper (II) oxide	<u>CuO</u>
9. gold (III) bromide	<u>AuBr₃</u>	19. lawrencium fluoride	<u>LrF₃</u>
10. zinc oxide	<u>ZnO</u>	20. actinium oxide	<u>Ac₂O₃</u>

B. Name the following compounds correctly. Only include the Roman Numeral in the name if the first element has 2 or more combining capacities.

1. CaBr ₂	<u>Calcium Bromide</u>	9. AsCl ₅	<u>Arsonic (V) Chloride</u>
2. K ₂ O	<u>Potassium Oxide</u>	10. Eu ₂ O ₃	<u>Europium (III) Oxide</u>
3. Fe ₂ O ₃	<u>Iron (III) Oxide</u>	11. NdBr ₃	<u>Neodymium Bromide</u>
4. ZnF ₂	<u>Zinc Fluoride</u>	12. Cu ₃ N ₂	<u>Copper (II) Nitride</u>
5. CuCl	<u>Copper (I) Chloride</u>	13. CoS	<u>Cobalt (II) Sulphide</u>
6. Ni ₂ O ₃	<u>Nickel (III) Oxide</u>	14. ZrO ₂	<u>Zirconium Oxide</u>
7. AuCl	<u>Gold (I) Chloride</u>	15. Li ₂ O	<u>Lithium Oxide</u>
8. HgBr	<u>Mercury (I) Bromide</u>	16. NaH	<u>Sodium Hydride</u>

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