

Name: _____ Course & Year: _____ Score: _____
Schedule: _____ Date Performed: _____ Date Submitted: _____
Laboratory Instructor: _____ Group Number: _____

Activity No. 4
WORKSHEET

CALORIMETRY AND SPECIFIC HEAT

Data, results, and calculations: Complete the table below, assume that the metal used is iron

(a) Weight of the calorimeter (inner vessel)	9.15 g
(b) Weight of Calorimeter + Water	36.70 g
(c) Weight of water, (b) - (a)	
(d) Initial temperature of water, (t_i)	28 °C
(e) Initial temperature of aluminum calorimeter, (t_{cal})	
(f) Initial temperature of metal (t_{metal})	98.0 °C
(g) Weight of metal sample, m_{metal}	7.5 g
(h) Final temperature of water, (t_f)	32.5 °C
(i) Final temperature of calorimeter, (t_{cal})	
(j) Final temperature of metal	
(k) Heat absorbed by the calorimeter, q_{cal}	
(l) Heat absorbed by the water, q_w	
(m) Experimental specific heat of metal sample	
(n) Theoretical specific heat of metal sample	0.460 cal/g°C
(o) Percent Error	%

GUIDE QUESTIONS

1	What is the purpose of stirring the contents of the calorimeter before the final temperature is taken?	
2	Why is it required to tap the hot metal with a tissue paper before placing the hot metal in the calorimeter?	
3	Why should the transferring of the hot metal from the beaker to the calorimeter be done quickly?	

Answer Sheet For Calorimetry Specific Heat

Daniela Niemeyer

Answer Sheet For Calorimetry Specific Heat:

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