

Name: _____ Course & Year: _____ Score: _____
 Schedule: _____ Date Performed: _____ Date Submitted: _____
 Laboratory Instructor: _____ Group Number: _____

Activity No. 4
 WORKSHEET

CALORIMETRY AND SPECIFIC HEAT

Data, results, and calculations: Complete the table below, assume that the metal used is iron

(a) Weight of the calorimeter (inner vessel)	9.15 g
(b) Weight of Calorimeter + Water	36.70 g
(c) Weight of water, (b) - (a)	
(d) Initial temperature of water, $(t)_{\text{w}}$	28 °C
(e) Initial temperature of aluminum calorimeter, $(t)_{\text{cal}}$	
(f) Initial temperature of metal $(t)_{\text{metal}}$	98.0 °C
(g) Weight of metal sample, m_{metal}	7.5 g
(h) Final temperature of water, $(t)_{\text{w}}$	32.5 °C
(i) Final temperature of calorimeter, $(t)_{\text{cal}}$	
(j) Final temperature of metal	
(k) Heat absorbed by the calorimeter, q_{cal}	
(l) Heat absorbed by the water, q_{w}	
(m) Experimental specific heat of metal sample	
(n) Theoretical specific heat of metal sample	0.460 cal/g°C
(o) Percent Error	%

GUIDE QUESTIONS

1	What is the purpose of stirring the contents of the calorimeter before the final temperature is taken?	
2	Why is it required to tap the hot metal with a tissue paper before placing the hot metal in the calorimeter?	
3	Why should the transferring of the hot metal from the beaker to the calorimeter be done quickly?	

Answer Sheet For Calorimetry Specific Heat

Daniela Niemeyer



Answer Sheet For Calorimetry Specific Heat:

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Table of Contents Answer Sheet For Calorimetry Specific Heat

1. Understanding the eBook Answer Sheet For Calorimetry Specific Heat
 - The Rise of Digital Reading Answer Sheet For Calorimetry Specific Heat
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Sheet For Calorimetry Specific Heat
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Sheet For Calorimetry Specific Heat
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Sheet For Calorimetry Specific Heat
 - Personalized Recommendations
 - Answer Sheet For Calorimetry Specific Heat User Reviews and Ratings
 - Answer Sheet For Calorimetry Specific Heat and Bestseller Lists
5. Accessing Answer Sheet For Calorimetry Specific Heat Free and Paid eBooks

- Answer Sheet For Calorimetry Specific Heat Public Domain eBooks
 - Answer Sheet For Calorimetry Specific Heat eBook Subscription Services
 - Answer Sheet For Calorimetry Specific Heat Budget-Friendly Options
6. Navigating Answer Sheet For Calorimetry Specific Heat eBook Formats
- ePub, PDF, MOBI, and More
 - Answer Sheet For Calorimetry Specific Heat Compatibility with Devices
 - Answer Sheet For Calorimetry Specific Heat Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Answer Sheet For Calorimetry Specific Heat
 - Highlighting and Note-Taking Answer Sheet For Calorimetry Specific Heat
 - Interactive Elements Answer Sheet For Calorimetry Specific Heat
8. Staying Engaged with Answer Sheet For Calorimetry Specific Heat
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Sheet For Calorimetry Specific Heat
9. Balancing eBooks and Physical Books Answer Sheet For Calorimetry Specific Heat
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Sheet For Calorimetry Specific Heat
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Sheet For Calorimetry Specific Heat
- Setting Reading Goals Answer Sheet For Calorimetry Specific Heat
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Sheet For Calorimetry Specific Heat
- Fact-Checking eBook Content of Answer Sheet For Calorimetry Specific Heat
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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