

Name: Samantha Norris Date: April 31 2020

## Student Exploration: Calorimetry Lab

**Vocabulary:** calorie, calorimeter, joule, specific heat capacity

### Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance's **specific heat capacity**. You will use the *Calorimetry Lab* Gizmo to determine the specific heat capacities of various substances.



1. On the **SIMULATION** pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the **GRAPH** tab, and click **Play** (▶).

A. What was the **Final temperature** of the copper and the water?

34.96

B. How much did the temperature of the copper change?

55.04

C. How much did the temperature of the water change?

4.96

2. Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

Water has a greater specific heat because water resists temperature better than copper.

# Answer Sheet For Calorimetry Lab

**Yasuhiko Syono, Murli H. Manghnani**

## **Answer Sheet For Calorimetry Lab:**

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