

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentration from pH and hydroxide ion concentration from pOH of a given solution.
3. Calculate the pH of a weak acid solution.

## Vocabulary

- |                   |                               |                     |
|-------------------|-------------------------------|---------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution    |
| 4. autoionization | 5. basic solution             | 6. neutral solution |
|                   |                               | 7. pH               |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the hydrogen ion concentration or hydroxide ion concentration of the solutions described. Round the concentration to the correct significant figures. Express values in scientific notation.

1. What substance is \_\_\_\_\_, contains hydrogen ions?
2. \_\_\_\_\_ is the concentration of those acids present in stomach acid equal to \_\_\_\_\_ molar.
3. \_\_\_\_\_ is the concentration of those acids present in \_\_\_\_\_ molar \_\_\_\_\_.
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# 19 2 Hydrogen Ions And Acidity Work Sheet Answers

**Wilhelm Mansfield Clark**



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