

## Chemistry Semester 1 Final Exam Review

1. What is a definition of Chemistry?  
**Study of matter**
  2. Matter includes all of the following EXCEPT:  
a. air      **b. light**      c. smoke      d. water vapor
  3. What is matter?  
**Something that has mass and takes up space (volume)**
  4. A physical change occurs when a  
a. peach spoils      c. copper bowl tarnishes  
b. bracelet turns your wrist green      **d. glue gun melts a glue stick**
  5. The particles in a solid are  
**a. packed closely together**      c. constantly in motion  
b. very far apart      d. able to slide past one another
- Of the following groups, 1, 2, 14, 15, 17, 18:
6. Group(s) **1, 2** contains only metals
  7. Group (s) **18** is called a Noble gases.
  8. Group(s) **17** has 7 electrons on the valence shell
  9. Group(s) **15** needs to gain 3 electrons to fill a stable octet
  10. Group(s) **2** has a 2+ oxidation state
  11. Group(s) **2** has an electron dot notation of  $\bullet X \bullet$
  12. Using the periodic table, an element similar to carbon would be **14**
  13. Based on their location on the periodic table, you could infer that **18/noble gases** are very unreactive
  14. The horizontal row on the periodic table is called a(n) **period**
  15. The vertical column of blocks on the periodic table is called a(n) **group/family**
  16. The elements on the zigzag line in the periodic table are **metalloids**
  17. Which is NOT a property of metal?  
a. malleability      c. unreactivity  
b. ability to lose electrons      d. ability to conduct heat
  18. A volume of 1 cc ( $\text{cm}^3$ ) is the same as **1** mL.
  19. To determine density, the quantities that must be measured are **volume** and **mass**
  20. The relationship between the mass  $m$  of a material, its volume  $V$ , and its density  $D$  is  
a.  $V = mD$       b.  $V/m = D$       **c.  $DF = m$**       d.  $D + V = m$
  28. The density of aluminum is  $2.70 \text{ g/cm}^3$ . The volume of a solid piece of Aluminum is  $1.50 \text{ cm}^3$ . Find the mass **4.05 grams**
  21. The density of pure diamond is  $3.5 \text{ g/cm}^3$ . The mass of a diamond is 0.25 g. Find the volume= **0.71  $\text{cm}^3$**
  22. What statement about density is true?  
a. two samples of a pure substance may have different densities if they are different sizes  
b. the density of a sample depends on its location on Earth  
c. a cylinder is always used to measure the volume  
**d. density is a physical property that remains constant for a pure substance**
  23. The number of grams equal to 0.5 kg is **500 g**
  24. In a graph the **independent** variable is placed on the x axis and the **dependant** variable is placed on the y axis.
  25. The variable that you change in the lab is the **independent** variable. The variable that responds to the change is called the **dependant** variable. All other things do not change and are called **constants**
  26. At first Dalton thought the atom was **solid, spherical, indivisible**, then Rutherford discovered the **nucleus** of the atom, then the subatomic particles **protons**, **neutrons**, and **electrons** were discovered
  27. The positively charged particle is the **proton** found in the **nucleus** of the atom
  28. The subatomic particle that has about the same mass as the proton, but with no electrical charge, is called a **neutron** and is found in the **nucleus** of the atom
  29. The **electron** is negatively charged and is found in the **electron cloud/outer shell** of the atom
  30. **Bohr** is the person credited with placing electrons on levels and **Schrodinger** established the electron cloud theory.
  31. The forces that hold the particles together in the nucleus are **columbic /electrostatic** forces.
  32. Isotopes are atoms of the same element that **have different number of neutrons**
  33. All isotopes of hydrogen contain **1 proton & 1 electron**
  34. What is the atomic number for Nitrogen? **7** To what subatomic particle does this correspond? **# of protons**
  35. Zinc -66 has **30** protons and **36** neutrons
  36. What is the electron configuration for sodium?  **$[\text{Ne}]3s^1$**  Chlorine?  **$[\text{Ne}]3s^23p^5$**  Neon?  **$[\text{He}]2s^22p^6$**
  37. What is the dot notation for magnesium? Bromine? Argon?

# Apexvs Spanish 1 Semester 2 Answer Key

**Robert R. Redfield**



## **Apexvs Spanish 1 Semester 2 Answer Key:**

## **Apexvs Spanish 1 Semester 2 Answer Key** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Apexvs Spanish 1 Semester 2 Answer Key**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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