

## AP Biology Study Guide Unit 1

### Properties of Water:

1. Polarity - allows H<sub>2</sub>O to form Hydrogen Bonds
2. Cohesion - Water H bonds to itself, creating surface tension
3. Adhesion - Water adheres to other polar substances
  - a. Allows easy transport of water against Gravity
4. Universal Solvent - Water can dissolve Polar and Ionic Compounds
  - a. This allows for reactions between Aqueous Solutions
5. Expands when Frozen - this means that Ice floats
6. High Specific Heat - water can absorb or release a large amount of heat with only a slight change in its own temperature (caused by cohesion)
  - a. High Heat of Vaporization - only warmest molecules can evaporate, so this causes evaporative cooling
7. Can be regulated by non-polar hydrophobic substances
8. Neutral PH → can form acids or bases if substances are added

Polarity - Difference in electronegativity between atoms bonded together create slightly charged ends of the molecules they form

Hydrophobic Substance - Nonpolar substance that repels water

Hydrophilic Substance - Polar substance that attracts water molecules

### CHON - Carbon, Hydrogen, Nitrogen, Oxygen

- Makeup 99% of living matter
- Strong Biological Cornerstones Because:
  - Smallest Diameter/Valency for atoms of 1, 2, 3, 4 valence electrons
  - Form Stable Covalent Bonds
  - Can form double and triple bonds
  - Versatile

### Biological Transition Metals - Mn, Fe, Co, Cu, Zn

- Can exist in multiple oxidation states → can be both oxidized and reduced

### Biological Ions: Na<sup>+</sup>, Mg<sup>2+</sup>, Cl<sup>-</sup>, K<sup>+</sup>, Ca<sup>2+</sup>

- Used to maintain homeostasis and osmotic concentration

Phosphorous/Sulphur - Energy Carriers which form strong high energy bonds

### 5 Foundation Biological Molecules:

# Apex Bio Study Guide

**Huangqi Zhang**

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## **Table of Contents Apex Bio Study Guide**

1. Understanding the eBook Apex Bio Study Guide
  - The Rise of Digital Reading Apex Bio Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Apex Bio Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Apex Bio Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Apex Bio Study Guide
  - Personalized Recommendations
  - Apex Bio Study Guide User Reviews and Ratings
  - Apex Bio Study Guide and Bestseller Lists

5. Accessing Apex Bio Study Guide Free and Paid eBooks
  - Apex Bio Study Guide Public Domain eBooks
  - Apex Bio Study Guide eBook Subscription Services
  - Apex Bio Study Guide Budget-Friendly Options
6. Navigating Apex Bio Study Guide eBook Formats
  - ePUB, PDF, MOBI, and More
  - Apex Bio Study Guide Compatibility with Devices
  - Apex Bio Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Apex Bio Study Guide
  - Highlighting and Note-Taking Apex Bio Study Guide
  - Interactive Elements Apex Bio Study Guide
8. Staying Engaged with Apex Bio Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Apex Bio Study Guide
9. Balancing eBooks and Physical Books Apex Bio Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Apex Bio Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Apex Bio Study Guide
  - Setting Reading Goals Apex Bio Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Apex Bio Study Guide
  - Fact-Checking eBook Content of Apex Bio Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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