

Unit 9 Progress Check: MCQ

Table 1

Cell Type	Operating Cell Potential for Commercial Batteries, E° (V)
Lithium-iodine	+2.80
Zinc-mercury	+1.35

Table 2

Half-Reaction	Standard Reduction Potential, E° (V)
$[\text{Zn}(\text{OH})_4]^{2-} + 2 e^- \rightarrow \text{Zn} + 4 \text{OH}^-$	-1.20
$\text{Zn}(\text{OH})_2 + 2 e^- \rightarrow \text{Zn} + 2 \text{OH}^-$	-1.25
$\text{HgO} + \text{H}_2\text{O} + 2 e^- \rightarrow \text{Hg} + 2 \text{OH}^-$	+0.30
$\text{O}_2 + 2 \text{H}_2\text{O} + 4 e^- \rightarrow 4 \text{OH}^-$	+0.40

Pacemakers are electronic devices that help regulate the heart rate. Currently, lithium-iodine cells are commonly used to power pacemakers and have replaced zinc-mercury cells. Table 1 provides the operating cell potential, E° , for each cell. Table 2 provides the standard reduction potentials for several half-reactions related to zinc-mercury and zinc-air cells.

1. Based on the information given, which of the following is a major difference between the zinc-mercury cell and the lithium-iodine cell?

(A) During the initial cell operation, each reaction is thermodynamically favorable, but the larger operating potential of the lithium-iodine cell indicates that its cell reaction is less thermodynamically favorable.

(B) During the initial cell operation, each reaction is thermodynamically favorable, but the larger operating potential of the lithium-iodine cell indicates that its cell reaction is more thermodynamically favorable.

(C) During the initial cell operation, the oxidation of iodine is thermodynamically favorable but the oxidation of mercury is not.

(D) During the initial cell operation, the oxidation of mercury is thermodynamically favorable but not the oxidation of iodine is not.

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