

Unit 9 Progress Check: MCQ

Table 1

Cell Type	Operating Cell Potential for Commercial Batteries, E (V)
Lithium-iodine	+2.80
Zinc-mercury	+1.35

Table 2

Half-Reaction	Standard Reduction Potential, E° (V)
$[\text{Zn}(\text{OH})_4]^{2-} + 2e^- \rightarrow \text{Zn} + 4\text{OH}^-$	-1.20
$\text{Zn}(\text{OH})_2 + 2e^- \rightarrow \text{Zn} + 2\text{OH}^-$	-1.25
$\text{HgO} + \text{H}_2\text{O} + 2e^- \rightarrow \text{Hg} + 2\text{OH}^-$	+0.10
$\text{O}_2 + 2\text{H}_2\text{O} + 4e^- \rightarrow 4\text{OH}^-$	+0.40

Pacemakers are electronic devices that help regulate the heart rate. Currently, lithium-iodine cells are commonly used to power pacemakers and have replaced zinc-mercury cells. Table 1 provides the operating cell potential, E , for each cell. Table 2 provides the standard reduction potentials for several half-reactions related to zinc-mercury and zinc-air cells.

1. Based on the information given, which of the following is a major difference between the zinc-mercury cell and the lithium-iodine cell?
- (A) During the initial cell operation, each reaction is thermodynamically favorable, but the larger operating potential of the lithium-iodine cell indicates that its cell reaction is less thermodynamically favorable.
- (B) During the initial cell operation, each reaction is thermodynamically favorable, but the larger operating potential of the lithium-iodine cell indicates that its cell reaction is more thermodynamically favorable. ✓
- (C) During the initial cell operation, the oxidation of iodine is thermodynamically favorable but the oxidation of mercury is not.
- (D) During the initial cell operation, the oxidation of mercury is thermodynamically favorable but not the oxidation of iodine is not.

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