

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
potassium-40 undergoes beta decay (electron)  
uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



### NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride  $\text{SF}_6$
- b. Lithium nitride  $\text{Li}_3\text{N}$
- c. Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride  $\text{SnCl}_2$
- e. Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )
- f. Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$
- g. Potassium bromate  $\text{KBrO}_3$
- h. Hydrosulfuric acid  $\text{H}_2\text{S(aq)}$
- i. Chloric acid  $\text{HClO}_3$
- j. Sulfurous acid  $\text{H}_2\text{SO}_3$

Name each of the following compounds.

- a.  $\text{CuSO}_4$  Copper (II) sulfate
- b.  $\text{AlF}_3$  aluminum fluoride
- c.  $\text{HI}$  hydrogen iodide
- d.  $\text{NO}$  nitrogen monoxide
- e.  $\text{H}_2\text{Se(aq)}$  hydroselenic acid
- f.  $\text{HNO}_3$  nitric acid
- g.  $\text{HNO}_2$  nitrous acid
- h.  $\text{NaHSO}_4$  sodium hydrogen sulfate

### Half-life

Tritium ( $\text{H-3}$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



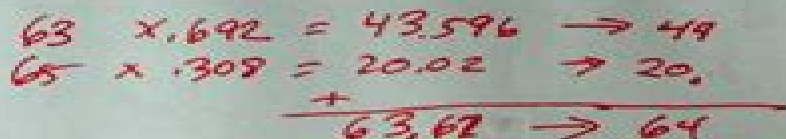
Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



# Answers Chapter 9 Mixed Review Modern Chemistry

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aqueous solutions of proteins or systems that are far from normal conditions e.g. critical or near critical mixtures. Second, the conventional thermodynamic methods are not sufficient for the accurate treatment of these mixtures. Last but not least, these systems are of interest for the pharmaceutical, biomedical and related industries. In order to meet the thermodynamic challenges involved in these complex mixtures, we employed a variety of traditional methods but also new methods such as the fluctuation theory of Kirkwood and Buff and *ab initio* quantum mechanical techniques. The Kirkwood-Buff (KB) theory is a rigorous formalism which is free of any of the approximations usually used in the thermodynamic treatment of multicomponent systems. This theory appears to be very fruitful when applied to the above-mentioned difficult systems. Applied Mechanics Reviews, 1989      **The Electrician**, 1897      **Curriculum Review**, 1982      **The Electrical Journal**, 1897      *The Al<sub>2</sub>SiO<sub>5</sub> Polymorphs* Derrill M. Kerrick, 2018-12-17. Volume 22 of *Reviews in Mineralogy* is a book exclusively devoted to three minerals: the Al<sub>2</sub>SiO<sub>5</sub> polymorphs andalusite, sillimanite and kyanite. It covers a wide variety of topics that must be considered in the analysis of metamorphic systems and illustrates the fact that modern metamorphic petrology demands an awareness of a wide spectrum of geologic variables and processes. This volume is intended to provide a comprehensive review summarizing the methods, theories and pitfalls of the various contributions on the aluminum silicates. Several sections of this book present the chronological development of research on various topics, giving readers historical perspectives on the development of theories, models and biases on various problems regarding the aluminum silicates. **Modern Powder Diffraction** David L. Bish, 1989. Principles of powder diffraction. Scattering by points. Calculation of the three-dimensional diffraction pattern. Instrumentation. Measurement of X-ray powder diffraction patterns: camera methods. Effect of instrument parameters. X-ray detectors. Techniques for the selection of monochromatic radiation. Resolutions versus intensity. Experimental procedures. Factors influencing d-spacing accuracy. Factors influencing intensity accuracy. Automation of qualitative phase analysis. Use of standards. Sample preparation for X-ray diffraction. Randomly oriented powder preparation. Preparing oriented clay mineral aggregates. Quantitative analysis. Absorption diffraction method. Example use of the reference intensity ratio or Rietveld method. Diffraction by small and disordered crystals. Mixed-layered micas and clay minerals. Diffraction from glass. Computer analysis of diffraction data. Profile fitting of powder diffraction patterns. Rietveld refinement of crystal structures using powder X-ray diffraction data. Synchrotron powder diffraction. Neutron powder diffraction. Reaction neutron sources. Special sample environments in neutron diffraction. Pollution Control Review, 1972      Civil Engineering and Public Works Review, 1969      **Mathematical Reviews**, 2003      *British Medical Journal*, 1897      *Collected Readings in Inorganic Chemistry* Gordon Galloway, 1972

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