

7. At a high temperature, 0.50 mol of HBr was placed in a 1.0 L container and allowed to decompose according to the reaction:



At equilibrium the $[\text{Br}_2]$ was measured to be 0.13 M. What is K_{eq} for this reaction at this temperature?



| | | | |
|-----|-------|-------|-------|
| [I] | 0.50 | 0 | 0 |
| [C] | -0.26 | +0.13 | +0.13 |
| [E] | 0.24 | 0.13 | 0.13 |

$$K_{\text{eq}} = \frac{[\text{H}_2][\text{Br}_2]}{[\text{HBr}]^2} = \frac{(0.13)^2}{(0.24)^2} = 0.29$$

Answer $K_{\text{eq}} = 0.29$

8. When 1.0 mol of $\text{NH}_3\text{(g)}$ and 0.40 mol of $\text{N}_2\text{(g)}$ are placed in a 5.0 L vessel and allowed to reach equilibrium at a certain temperature, it is found that 0.78 mol of NH_3 is present. The reaction is:



$$\text{initial } [\text{NH}_3] = \frac{1.0 \text{ mol}}{5.0 \text{ L}} = 0.20 \text{ M} \quad \text{initial } [\text{N}_2] = \frac{0.40 \text{ mol}}{5.0 \text{ L}} = 0.080 \text{ M}$$

$$\text{equil. } [\text{NH}_3] = \frac{0.78 \text{ mol}}{5.0 \text{ L}} = 0.156 \text{ M}$$



| | | | |
|-----|--------|--------|--------|
| [I] | 0.20 | 0 | 0.080 |
| [C] | -0.044 | +0.066 | +0.012 |
| [E] | 0.156 | 0.066 | 0.102 |

a) Calculate the equilibrium concentrations of all three species.

$$[\text{NH}_3] = 0.156 \text{ M} \quad [\text{H}_2] = 0.066 \text{ M} \quad [\text{N}_2] = 0.102 \text{ M}$$

b) Calculate the value of the equilibrium constant at this temperature.

(use unrounded concs, then round to 3 sig figs)

$$K_{\text{eq}} = \frac{[\text{H}_2]^3 [\text{N}_2]}{[\text{NH}_3]^2} = \frac{(0.066)^3 (0.102)}{(0.156)^2} = 0.0012$$

Answer 1.2×10^{-3}

c) How many moles of H_2 are present at equilibrium?

$$0.066 \text{ M} \times 5.0 \text{ L} = 0.33 \text{ mol}$$

Answer 0.33 mol

d) How many moles of N_2 are present at equilibrium?

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