



## General Chemistry II Equation Sheet

Think about how to set up the problem first, then apply the needed principles and formulas.

### Phase Changes

$$\Delta T_b = iK_b m$$

$$\Delta T_f = iK_f m$$

$$\pi = iMRT$$

$$c = kP$$

$$\frac{c_1}{P_1} = \frac{c_2}{P_2}$$

$$P_{\text{solution}} = \chi_{\text{solvent}} P_{\text{solvent}}$$

$$\Delta P = P_{\text{solvent}} - P_{\text{solution}} = \chi_{\text{solute}} P_{\text{solvent}}$$

### Cells

$$E = E^\circ - \frac{RT}{nF} \ln(Q)$$

$$E = E^\circ - \frac{0.0592}{n} \log(Q) \text{ (at } 25^\circ\text{C)}$$

$$\Delta G = -nFE_{\text{cell}}$$

$$\Delta G^\circ = -nFE_{\text{cell}}^\circ$$

$$E_{\text{cell}}^\circ = \frac{RT}{nF} \ln(K)$$

$$E_{\text{cell}}^\circ = \frac{0.0592V}{n} \log(K) \text{ (at } 25^\circ\text{C)}$$

USE ONE:

$$E_{\text{cell}}^\circ = E_{\text{cathode(red.)}}^\circ - E_{\text{anode(red.)}}^\circ$$

$$E_{\text{cell}}^\circ = E_{\text{cathode(red.)}}^\circ + E_{\text{anode(ox.)}}^\circ$$

### Thermodynamics

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta S_{\text{pure}} = -\frac{\Delta H_{\text{pure}}}{T}$$

$$\Delta S_{\text{univ}} = \Delta S_{\text{sys}} + \Delta S_{\text{surv}}$$

$$\Delta S_{\text{univ}}^\circ = \sum S_{\text{products}}^\circ + \sum S_{\text{reactants}}^\circ$$

$$\Delta G_{\text{univ}}^\circ = \sum \Delta G_{\text{products}}^\circ + \sum \Delta G_{\text{reactants}}^\circ$$

### Concentrations

$$\text{molality} = \frac{\text{moles solute}}{\text{moles solvent kg}}$$

$$\% \text{ by mass} = \frac{\text{moles solute}}{\text{moles solute} + \text{moles solvent}} \times 100\%$$

$$\text{parts per million} = \frac{\text{moles solute}}{\text{moles solute} + \text{moles solvent}} \times 10^6$$

$$\text{parts per billion} = \frac{\text{moles solute}}{\text{moles solute} + \text{moles solvent}} \times 10^9$$

### Acids and Bases

$$K_w = [H_3O^+][OH^-] = K_a \times K_b$$

$$pOH = -\log[OH^-]$$

$$[OH^-] = 10^{-pOH}$$

$$pH + pOH = pK_w + pK_b = 14.00 \text{ (at } 25^\circ\text{C)}$$

$$pH = pK_w + \log \frac{[A^-]}{[HA]}$$

$$pOH = pK_b + \log \frac{[BH^+]}{[B]}$$

$$\% \text{ ionization} \frac{[H_3O^+]_{\text{eq}}}{[H_3O^+]_0} \times 100\%$$

$$pK_w = -\log(K_w)$$

$$pK_b = -\log(K_b)$$

$$K_w = 10^{-pK_w}$$

$$K_b = 10^{-pK_b}$$

### Equilibria

$$Q_c = \frac{[C]^x [D]^y}{[A]^a [B]^b}$$

$$K_c = \frac{[C]^x [D]^y}{[A]^a [B]^b}$$

$$Q_p = \frac{P_C^x P_D^y}{P_A^a P_B^b}$$

$$K_p = \frac{P_C^x P_D^y}{P_A^a P_B^b}$$

$$K_p = K_c [R \times T]^{\Delta n}$$

$$\Delta G = \Delta G^\circ + RT \ln(Q)$$

$$\Delta G^\circ = -RT \ln(K)$$

$$\text{rate} = k[A]^x [B]^y$$

$$k = Ae^{-\frac{E_a}{RT}}$$

$$\ln\left(\frac{P_2}{P_1}\right) = \frac{E_a}{RT} \left(\frac{1}{T_2} - \frac{1}{T_1}\right)$$

$$\text{half life} = \frac{\ln(2)}{E_a} \text{ (1st order reactions only)}$$

$$[A]_t = -kt + [A]_0$$

$$\ln([A]_t) = -kt + \ln([A]_0)$$

$$\frac{1}{[A]_t} = kt + \frac{1}{[A]_0}$$



# Answer Sheet General Chemistry Equations

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