

CHEMISTRY ADVANCED STUDY

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FUNDAMENTAL CONCEPTS

Math

- Properties of exponents
- Properties of logarithms
 - $\log_b(M^pN^q) = \log_b(M) + \log_b(N)$
 - $\log_b(MN) = \log_b(M) \cdot \log_b(N)$
 - $\log_b(M^p) = p \log_b(M)$
- Graphing slope integrals.
 - $y = mx + b$
 - b is the y -intercept, m is the slope
- Quadratic formula
 - $x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$
 - Ex. $x^2 + 4x - 21 = 0$
 $x = \frac{-4 \pm \sqrt{16 - 4(1)(-21)}}{2}$
 - $x = -2 \pm 5$
 - $x = 3$ or $x = -7$

Science

- Scientific notation
- Dimensional analysis
- Significant figures

Physics and Biology

- Scientific method
- Coulomb's law: electric force between like or opposite charges; dependence on separation distance and charge magnitude

ELEMENTS AND ATOMS

Protons: define the element (atomic number), positive

Electrons: define the charge (bonds and polarity), negative

Neutrons: define isotopes, neutral

ATOMIC MASS AND WEIGHT

Atomic Mass: Carbon-12 (exactly 12 atomic mass unit)

Atomic Weight

- weighted average of atomic masses of element
- $(\text{atomic mass} \cdot \text{neutrons}) + (\text{atomic mass} \cdot \text{protons})$

MOLE A.K.A. AVOGADRO'S NUMBER

$\approx 6.02 \cdot 10^{23}$

In Terms of Carbon

1 mole of ^{12}C atoms in 12g of Carbon

$6.02 \cdot 10^{23} = 12\text{g of Carbon}$

Gram to AMU

1 gram = 1 mole of atomic mass units

1 gram = $6.02 \cdot 10^{23}$ amu

1 mol ^{27}Al = 27g

2 mol ^{56}Fe = 56g

4 mol ^{16}O = 64g

Atoms And Moles Study Guide

O García

Atoms And Moles Study Guide:

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