

**Key Concept Summary**

# Atoms, Bonding, and the Periodic Table

## What Determines an Element's Chemistry?

An atom's electrons are found in different energy levels. Valence electrons have the highest energy. Each element has a specific number of valence electrons, from 1 to 8. The number of valence electrons in the atoms of an element determines that element's chemical properties.

An electron dot diagram includes the symbol for an element surrounded by dots. Each dot represents one valence electron. Atoms with 8 valence electrons tend to be more stable, so in forming compounds, atoms tend to obtain 8 valence electrons. Valence electrons may be transferred or shared. A chemical bond is the force of attraction that holds 2 atoms together as a result of the arrangement of electrons between them.

The periodic table includes rows, called periods, and columns, called groups. The atomic number increases from left to right across each period, as does the number of valence electrons. The first element in each period has 1 valence electron. The last element (except Period 1) has 8 valence electrons. Because elements within a group (except Period 1) have the same number of valence electrons, they have similar properties.

Except for helium, each of the Group 18 elements, the noble gases, has 8 valence electrons. They are stable, unlikely to gain, lose, or share electrons, and do not react easily with other elements. Metal atoms react by losing their valence electrons. A metal's reactivity depends on how easily it loses valence electrons. In general, reactivity decreases from left to right across the periodic table. Nonmetals become stable when their atoms gain or share enough electrons to have 8 valence electrons. Nonmetals usually combine with metals by gaining electrons, but can also combine with other nonmetals and metalloids by sharing. The metalloids, which lie between metals and nonmetals in the periodic table, can either lose or share electrons. They have some properties of both metals and nonmetals.

Hydrogen is placed in Group 1 because it has one valence electron, but it is considered a nonmetal. Its properties differ greatly from those of the other elements in Group 1. When it combines with other nonmetals, hydrogen shares its electron and forms a stable arrangement of 2 electrons.

**On a separate sheet of paper, use the periodic table to create electron dot diagrams for barium (Ba), carbon (C), cesium (Cs), iodine (I), krypton (Kr), phosphorous (P), and strontium (Sr).**

# Atoms And Bonding Review And Reinforce Answers

**C.B. Duke,E. Ward Plummer**

## Atoms And Bonding Review And Reinforce Answers:

**AP Biology** Deborah T. Goldberg, 2020-03-03 Barron's AP Biology With Two Practice Tests is revised to reflect all upcoming changes to the AP Biology course and the May 2020 exam. You'll get the in-depth content review and practice tests you need to fully prepare for the exam. This edition features Two full length practice exams in the book that follow the content and style of the revised AP Biology exam with detailed answer explanations for all questions. A fully revised introduction that covers the new exam format including the exam sections, the question types, the number of questions per section and the amount of time allotted per section. Helpful test taking tips and strategies throughout the book plus icons that designate sections with particularly helpful background information to know. 19 comprehensive review chapters that cover all of the major topic areas that will be tested on the exam including the Cell Cycle, Photosynthesis, Heredity and much more. End of chapter practice questions that reinforce the concepts reviewed in each chapter. Appendices with key measurements that you should be familiar with as well as a glossary of key terms and definitions.

*Prentice Hall Chemistry*, 2000

**AP Biology Premium** Deborah T. Goldberg, 2020-03-03 Barron's AP Biology is one of the most popular test preparation guides around and a must have manual for success on the Biology AP Test. In this updated book, test takers will find Two full length exams that follow the content and style of the new AP exam. All test questions answered and explained. An extensive review covering all AP test topics. Hundreds of additional multiple choice and free response practice questions with answer explanations. This manual can be purchased alone or with an optional CD ROM that includes two additional practice tests with answers and automatic scoring. BONUS ONLINE PRACTICE TEST Students who purchase this book or package will also get FREE access to one additional full length online AP Biology test with all questions answered and explained. Want to boost your studies with even more practice and in-depth review? Try Barron's Ultimate AP Biology for even more prep.

*Prentice Hall Science Explorer* Michael J. Padilla, 2002

*Environmental Science and Technology* Diana L. Turner, 2003

**Biology** Kenneth Raymond Miller, Prentice Hall (School Division), 1999-02

**Ess Chem Probs Study Guide** David Margolese, David W. Oxtoby, 2004-04

**Media Review**, 1986

**Biology** Gerhart Campbell, 2000-09

**Ecology**, 1994 Authorized teaching resource in Alberta for senior high science 14 24 1995 2004

**The Software Encyclopedia 2001**, 2001

**Chemical Interactions**, 2005

**Matter, Building Block of the Universe**, 1993

**Frontiers in Surface Science and Interface Science** C.B. Duke, E. Ward Plummer, 2002-05-21

Any notion that surface science is all about semiconductors and coatings is laid to rest by this encyclopedic publication. Bioengineered interfaces in medicine, interstellar dust, DNA computation, conducting polymers, the surfaces of atomic nuclei all are brought up to date. *Frontiers in Surface and Interface Science* a milestone publication deserving a wide readership. It combines a sweeping expert survey of research today with an educated look into the future. It is a future that embraces surface phenomena on scales from the subatomic to the galactic as well as traditional topics like semiconductor design, catalysis and surface processing modeling and characterization. And great efforts

have been made to express sophisticated ideas in an attractive and accessible way Nanotechnology surfaces for DNA computation polymer based electronics soft surfaces interstellar surface chemistry all feature in this comprehensive collection **An Introduction to Physical Science** James T. Shipman, Jerry D. Wilson, Aaron W. Todd, 2002-07-08 This text presents a survey of the physical sciences physics chemistry astronomy meteorology and geology for non science majors Topics are treated both descriptively and quantitatively providing flexibility for instructors who wish to emphasize a highly descriptive approach a highly quantitative approach or any spectrum in between *Previews*, 1978 *Aeronautical Engineering Review*, 1955 **The Software Encyclopedia 2000** Bowker Editorial Staff, 2000-05 *Sif Chemistry Nl Tb* Rex M. Heyworth, 2007 South African Engineer and Electrical Review, 1955

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