

Key Concept Summary

Atoms, Bonding, and the Periodic Table

What Determines an Element's Chemistry?

An atom's electrons are found in different energy levels. Valence electrons have the highest energy. Each element has a specific number of valence electrons, from 1 to 8. The number of valence electrons in the atoms of an element determines that element's chemical properties.

An electron dot diagram includes the symbol for an element surrounded by dots. Each dot represents one valence electron. Atoms with 8 valence electrons tend to be more stable, so in forming compounds, atoms tend to obtain 8 valence electrons. Valence electrons may be transferred or shared. A chemical bond is the force of attraction that holds 2 atoms together as a result of the arrangement of electrons between them.

The periodic table includes rows, called periods, and columns, called groups. The atomic number increases from left to right across each period, as does the number of valence electrons. The first element in each period has 1 valence electron. The last element (except Period 1) has 8 valence electrons. Because elements within a group (except Period 1) have the same number of valence electrons, they have similar properties.

Except for helium, each of the Group 18 elements, the noble gases, has 8 valence electrons. They are stable, unlikely to gain, lose, or share electrons, and do not react easily with other elements. Metal atoms react by losing their valence electrons. A metal's reactivity depends on how easily it loses valence electrons. In general, reactivity decreases from left to right across the periodic table. Nonmetals become stable when their atoms gain or share enough electrons to have 8 valence electrons. Nonmetals usually combine with metals by gaining electrons, but can also combine with other nonmetals and metalloids by sharing. The metalloids, which lie between metals and nonmetals in the periodic table, can either lose or share electrons. They have some properties of both metals and nonmetals.

Hydrogen is placed in Group 1 because it has one valence electron, but it is considered a nonmetal. Its properties differ greatly from those of the other elements in Group 1. When it combines with other nonmetals, hydrogen shares its electron and forms a stable arrangement of 2 electrons.

On a separate sheet of paper, use the periodic table to create electron dot diagrams for barium (Ba), carbon (C), cesium (Cs), iodine (I), krypton (Kr), phosphorous (P), and strontium (Sr).

Atoms And Bonding Review And Reinforce Answers

C.B. Duke,E. Ward Plummer



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