

## Atoms and Bonding

Multiple-Choice

<sup>10</sup> The other member of the Committee' was Dr. John C. St. John, who was also the first

1. A(n) \_\_\_\_\_ is an atom or group of atoms that has a net electric charge.  
a. compound  
b. ion  
c. crystal  
d. molecule

2. A covalent bond is called \_\_\_\_\_ if the atoms equally share one or more pairs of electrons.  
a. ionic  
b. ionic bond  
c. nonpolar  
d. polar

3. An attraction between a positive metal ion and the electrons surrounding it is a.  
a. covalent bond  
b. metallic bond  
c. nonpolar bond  
d. triple bond

4. Which of the following does NOT apply to valence electrons?  
a. loosely held by the nucleus  
b. shown on electron dot diagrams  
c. transferred or shared in bonding  
d. make up most of the atom's mass

5. The number of \_\_\_\_\_ remains the same from top to bottom in a group in the periodic table.  
a. protons  
b. valence electrons  
c. electrons  
d. neutrons

6. Elements in Group 17, the halogen family, are very reactive because their atoms have \_\_\_\_\_ valence electrons.  
a. zero  
b. one  
c. seven  
d. eight

7. A molecule in which the polar covalent bonds do not cancel out will be  
a. nonpolar  
b. polar  
c. ionic  
d. metallic

8. Which of the following is NOT a characteristic that results from metallic bonding?  
a. hard, brittle crystals  
b. electrical conductivity  
c. ability to be hammered into sheets  
d. ductile

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**Eric D. Leskowitz**

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