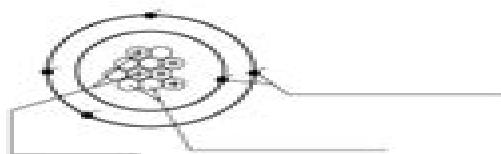


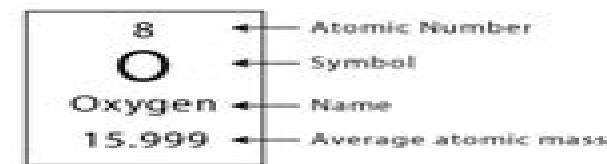
Basic Atomic Structure

1. Label the parts of an atom in the diagram below



- What type of charge does a proton have? _____
- What type of charge does a neutron have? _____
- What type of charge does an electron have? _____
- Which two subatomic particles are located in the nucleus of an atom? _____

2. An element is represented by its chemical symbol along with a few numbers. Based on the example on the right, fill in the numbers of protons (P), neutrons (N), and electrons (E) for the following elements.



6 C Carbon 12.011	#P _____	#N _____	#E _____
-----------------------------------	----------	----------	----------

10 Ne Neon 20.108	#P _____	#N _____	#E _____
-----------------------------------	----------	----------	----------

19 K Potassium 39.098	#P _____	#N _____	#E _____
---------------------------------------	----------	----------	----------

3. Complete the table using your knowledge of the periodic table.

Symbol	Atomic Number	Mass Number	Number of Protons	Number of Electrons	Number of Neutrons
Na			11		12
K		39		19	
			38		50
F				9	10
	20	40		20	
	50			50	69
I	53	127			

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Arthur P. Cracknell

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