

## Q3 Chapter 9 Preparation Problems: Chemical Reactions and Equations

### Section 9.10: Precipitation Reactions

1. Which of the changes in ionic composition would account for the following reaction?  
a. no change      b. oxidation      c. nuclear, more electrons lost  
d. reduction      e. reduction      f. ion exchange

2. How many electrons does the overall ionic group in  $\text{Fe}^{2+}$  have available for oxidation?  
a.  $\text{Cl}^{2-}$       b.  $\text{Br}^-$       c.  $\text{O}^{2-}$   
d.  $\text{F}^-$       e.  $\text{ClO}_4^-$       f.  $\text{NO}_3^-$

3. Which reaction(s) below would not produce an ion pair?

a.  $\text{Ba}^{2+}$       b.  $\text{H}^+$       c.  $\text{H}^+$   
d.  $\text{ClO}_4^-$       e.  $\text{H}^+$       f.  $\text{Na}^{+}$

4. Which of the following involving the charges for each ion,  $\text{Li}^{+}$  and  $\text{Cl}^{-}$ , are all incorrect?  
a. anions/oxides      b. cations/oxides      c. anions/oxides  
d. cations/oxides      e. hydronium ions      f. anions/oxides

5. Which of the following ions,  $\text{Al}^{3+}$ ,  $\text{Fe}^{2+}$ ,  $\text{Cl}^{-}$ ,  $\text{SO}_4^{2-}$ ,  $\text{H}^+$ , and  $\text{NO}_3^-$ , would not form an ion pair?  
a.  $\text{Cl}^{-}$       b.  $\text{Fe}^{2+}$       c.  $\text{Cl}^{-}$   
d.  $\text{SO}_4^{2-}$       e.  $\text{H}^+$       f.  $\text{NO}_3^-$

### 9.11: Precipitation and Filtering: Reactions that Involve $\text{CO}_3^{2-}$ Compounds

1. Write the formulas for these binary ionic compounds.  
a. hydrogen iodide      b. potassium iodide      c. sodium sulfide  
d. zinc hydroxide      e. aluminum iodide      f. silver bromide

2. Write the formulas for the corresponding binary ionic group given below.  
a.  $\text{Ba}^{2+}$  and  $\text{Cl}^{-}$       b.  $\text{Al}^{3+}$  and  $\text{O}^{2-}$       c.  $\text{Fe}^{2+}$  and  $\text{Cl}^{-}$   
d.  $\text{Mg}^{2+}$  and  $\text{I}^-$       e.  $\text{H}^+$  and  $\text{Br}^-$       f.  $\text{Fe}^{2+}$  and  $\text{O}^{2-}$

3. Name the following binary ionic compounds.  
a.  $\text{BaCl}_2$       b.  $\text{CaCl}_2$       c.  $\text{FeCl}_2$       d.  $\text{BaSO}_4$   
e.  $\text{BaOH}_2$       f.  $\text{BaI}_2$       g.  $\text{Ba}_2\text{O}$       h.  $\text{Ba}_2\text{S}_2\text{O}_3$

4. Write formulas for the following ionic compounds.  
a. sodium phosphate      b. sodium hydroxide      c. ammonium chloride  
d. ammonium sulfide      e. potassium phosphate      f. potassium chloride

5. Write formulas for the following ionic compounds.  
a.  $\text{FeCl}_3$  and  $\text{NaCl}_2$       b. barium ions and hydroxide ions  
d.  $\text{H}^+$  and  $\text{Na}^{+}$       e. barium ions and sulfide ions

# Answers Pearson Chemical Names And Formulas Packet

**S Ashworth**

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