

Question 3

(a) The Avogadro number is the number of carbon atoms in 12 g of carbon-12 and is numerically equal to 6.02×10^{23} .

$$(i) \quad \frac{1 \text{ mole}}{2.30} = \frac{6.02 \times 10^{23}}{3.01 \times 10^{22}}$$

molar mass

Cross multiply

$$\frac{2.30 \times 6.02 \times 10^{23}}{\text{molar mass}} = 3.01 \times 10^{22}$$

$$\begin{aligned} \text{molar mass of } x &= \frac{2.30 \times 6.02 \times 10^{23}}{3.01 \times 10^{22}} \\ &= 2.30 \times 2 \times 10 \\ &= 46.0 \text{ g/mol} \end{aligned}$$

$$(9) \quad \begin{aligned} N_x + O_y &= 46 \\ 14x + 16y &= 46 \end{aligned}$$

If $x=1$,

$$14 + 16y = 46$$

$$16y = 46 - 14 = 32$$

$$y = \frac{32}{16} = 2$$

NB: If $x=2$,
y gives decimal

Hence compound is NO_2

- (b) (i) CCl_4 is not soluble in water, hence nothing happens.
 (ii) SiCl_4 is decomposed by water to hydrochloric acid with evolution of heat.

(c) Silicon has empty 3d orbitals available to accept a lone pair from the water molecule. Carbon doesn't have empty 2-level orbitals. This means that the oxygen can bond to the silicon before the need to break a silicon-chlorine bond.

Answers Of Chemistry 2 Essay And 1 Obj

K Morrison



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