

### QUESTION 3

(a) The Avogadro number is the number of carbon atoms in 12 g of carbon-12 and is numerically equal to  $6.02 \times 10^{23}$ .

$$(i) \text{Molar mass} = \frac{6.02 \times 10^{23}}{2.30} = 3.01 \times 10^{22}$$

molar mass

Cross multiply -

$$\frac{2.30 \times 6.02 \times 10^{23}}{\text{molar mass}} = 3.01 \times 10^{22}$$

$$\text{molar mass of x} = \frac{2.30 \times 6.02 \times 10^{23}}{3.01 \times 10^{22}}$$

$$= 2.30 \times 2 \times 10^1$$

$$= 46.0 \text{ g/mol}$$

$$(ii) \begin{cases} \cancel{N_x} + O_y = 46 \\ 14x + 16y = 46 \end{cases}$$

IF  $x = 1$ ,

$$14 + 16y = 46$$

$$16y = 46 - 14 = 32$$

$$y = \frac{32}{16} = 2$$

NB: If  $x = 2$ ,  
y gives decimal

Hence compound is  $NO_2$ .

(b) (i)  $CCl_4$  is not soluble in water, hence nothing happens.

(ii)  $SiCl_4$  is decomposed by water to hydrochloric acid with evolution of heat.

(c) Silicon has empty 3d orbitals available to accept a lone pair from the water molecule. Carbon doesn't have empty 2s-level orbitals. This means that the oxygen can bond to the silicon before the need to break a silicon-chlorine bond.

# Answers Of Chemistry 2 Essay And 1 Obj

**K Morrison**

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