

2. Tertiary acids -- just like binary acids, tertiary acids must begin with hydrogen. These acids also contain a polyatomic ion.

Polyatomic ion	Name of ion	Polyatomic ion	Name of ion
$\text{OH}^-$	Hydroxide	$\text{PO}_4^{3-}$	Phosphate
$\text{O}_2^{2-}$	Peroxide	$\text{ClO}_4^-$	perchlorate
$\text{C}_2\text{O}_4^{2-}$	Acetate	$\text{ClO}_3^-$	chlorate
$\text{S}_2^{2-}$	Sulfide	$\text{ClO}_2^-$	chlorite
$\text{S}_2^{2-}$	sulfate	$\text{ClO}^-$	hypochlorite

Tertiary acids are three element acids. Most tertiary acids are OXYACIDS containing hydrogen, oxygen and other element.

Naming:

- Name the anion within the acid
- Change the anion ending to one of the following:

Anion ending	In acid, replace with:
-ide	-ous acid
-ate	-ic acid

Formula	anion	Tertiary compound name	Tertiary acid name
$\text{H}_2\text{SO}_3$	$\text{SO}_3^{2-}$	Hydrogen Sulfite	Sulfurous acid
$\text{H}_2\text{SO}_4$	$\text{SO}_4^{2-}$	Hydrogen sulfate	Sulfuric acid
$\text{HNO}_2$	$\text{NO}_2^-$	Hydrogen Nitrite	Nitrous acid
$\text{HNO}_3$	$\text{NO}_3^-$	Hydrogen Nitrate	Nitric acid
$\text{HClO}_4$	$\text{ClO}_4^-$	perchlorate	Perchloric acid
$\text{HClO}_3$	$\text{ClO}_3^-$	chlorate	Chloric acid
$\text{HClO}_2$	$\text{ClO}_2^-$	chlorite	Chlorous acid
$\text{HClO}$	$\text{ClO}^-$	hypochlorite	Hypochlorous acid

# Answers For Writing Ternary Formulas

**Clifford Mitchell**



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