

Name _____ Date _____ Class _____

CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

States of Matter

Section 13.1 Gases

In your textbook, read about the kinetic molecular theory.

Complete each statement.

1. The kinetic molecular theory describes the behavior of gases in terms of particles in motion.
2. The kinetic molecular theory makes the following assumptions:
 - a. In a sample of a gas, the volume of the gas particles themselves is very small compared to the volume of the sample.
 - b. Because gas particles are far apart, there are no significant attractive or repulsive forces between gas particles.
 - c. Gas particles are in constant and random motion.
 - d. The collisions between gas particles are elastic; that is, no kinetic energy is lost.
3. The kinetic energy of a particle is represented by the equation $KE = \frac{1}{2}mv^2$.
4. Temperature is a measure of the average kinetic energy of the particles in a sample of matter.

In your textbook, read about explaining the behavior of gases.

For each statement below, write true or false.

- | | |
|--------------|--|
| <u>True</u> | 1. Gases are less dense than solids because there is a lot of space between the particles of a gas. |
| <u>True</u> | 2. The random motion of gas particles causes a gas to expand until it fills its container. |
| <u>False</u> | 3. The density of a gas decreases as it is compressed. |
| <u>True</u> | 4. A gas can flow into a space occupied by another gas. |
| <u>True</u> | 5. The diffusion of a gas is caused by the random motion of the particles of the gas. |
| <u>False</u> | 6. Lighter gas particles diffuse less rapidly than do heavier gas particles. |
| <u>True</u> | 7. During effusion, a gas escapes through a tiny opening into a vacuum. |
| <u>False</u> | 8. Graham's law of effusion states that the rate of effusion for a gas is directly related to the square root of its molar mass. |

Name _____ Date _____ Class _____

CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

Section 13.2 continued

In your textbook, read about gas pressure.

Circle the letter of the choice that best completes the statement or answers the question.

13. Pressure is defined as force per unit c. area.
14. What is an instrument designed to measure atmospheric pressure? c. sphygmomanometer
15. The height of the liquid in a barometer is affected by all of the following EXCEPT the a. altitude.
16. The pressure of the gas in a manometer is directly related to which of the following quantities? b. height of the mercury column in the closed end arm
17. One atmosphere is equal to a pressure of b. 101.3 kPa.
18. The partial pressure of a gas depends on all of the following EXCEPT the b. size of the container.
19. The pressure of a sample of air in a manometer is 102.1 kPa. What is the partial pressure of nitrogen (N_2) in the sample if the combined partial pressures of the other gases is 22.4 kPa? b. 79.7 kPa

Use the figure to answer the following questions.

20. What instrument is illustrated in the figure? barometer
21. Who invented this instrument? Evangelista Torricelli
22. What are the two opposing forces that control the height of the mercury in the column? gravity and atmospheric pressure
23. What does it mean when the level of mercury drops in the column? The atmospheric pressure has increased.



Answers For Study Guide States Of Matter

LL Leslie



Answers For Study Guide States Of Matter:

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Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry **Class 3**

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Table of Contents Answers For Study Guide States Of Matter

1. Understanding the eBook Answers For Study Guide States Of Matter
 - The Rise of Digital Reading Answers For Study Guide States Of Matter
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Study Guide States Of Matter
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Study Guide States Of Matter
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Study Guide States Of Matter
 - Personalized Recommendations
 - Answers For Study Guide States Of Matter User Reviews and Ratings
 - Answers For Study Guide States Of Matter and Bestseller Lists

5. Accessing Answers For Study Guide States Of Matter Free and Paid eBooks
 - Answers For Study Guide States Of Matter Public Domain eBooks
 - Answers For Study Guide States Of Matter eBook Subscription Services
 - Answers For Study Guide States Of Matter Budget-Friendly Options
6. Navigating Answers For Study Guide States Of Matter eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Study Guide States Of Matter Compatibility with Devices
 - Answers For Study Guide States Of Matter Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Study Guide States Of Matter
 - Highlighting and Note-Taking Answers For Study Guide States Of Matter
 - Interactive Elements Answers For Study Guide States Of Matter
8. Staying Engaged with Answers For Study Guide States Of Matter
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Study Guide States Of Matter
9. Balancing eBooks and Physical Books Answers For Study Guide States Of Matter
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Study Guide States Of Matter
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers For Study Guide States Of Matter
 - Setting Reading Goals Answers For Study Guide States Of Matter
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers For Study Guide States Of Matter
 - Fact-Checking eBook Content of Answers For Study Guide States Of Matter
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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