

Rxnction Energy

QUESTION 1

Consider the following following conditions for the reaction given below:

- The reaction is the same overall order, but $k_{\text{cat}} = 100$ s⁻¹.
- The activation energy of the reaction is the same as the activation energy for the following elementary step:



- Activation energy for the reaction is the same as the activation energy for the following elementary step:
- What is the apparent rate order for the reaction?

Consider the following conditions for the reaction given below:

- At 298 K, the rate of the reaction is 1000 times faster than at 273 K. What is the activation energy for the reaction?

- Activation energy for the reaction is 1000 times faster than at 273 K. What is the rate of the reaction at 298 K?

- The pre-exponential factor for the reaction is 1000, and the activation barrier is constant at 1000 J/mol.

- What is the order of the reaction?

Answers For Reaction Energy Section 1

M Mark

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