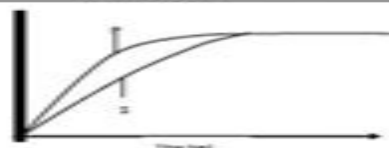


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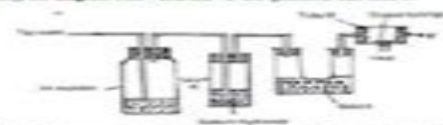
Mock Exams + Answers



- (ii) Which of the two curves represents the reaction of 2M concentrated HCl with powdered calcium carbonate. Give a reason. (2 marks)
- (iii) Why do the two curves flatten at the same level of production of CO_2 ? (1 mark)
4. Study the following equilibrium reaction.
 $\text{CO(g)} + \text{H}_2\text{O(g)} \rightleftharpoons \text{CO}_2\text{(g)} + \text{H}_2\text{(g)}$ (1 mark)
- Suggest two ways of increasing the yield of CO_2 . (2 marks)
7. The table below gives some elements in the periodic table. Use it to answer the questions that follow. The letters do not represent the actual symbols of the elements.
- | Element | A | B | C | D | E |
|---------------|----|----|----|----|----|
| Atomic number | 12 | 15 | 14 | 13 | 16 |
- (i) A metallic element which forms ions with the smallest ionic radius? Explain. (1 mark) (ii) A non-metallic element with the largest atomic size? Explain. (1 mark)



- (iii) Write the equation for the reaction taking place in region A. (1 mark)
- (iv) Name the gas produced at region B. (1 mark)
- (v) State ONE use of the gas named in (iv) above. (1 mark)



- (i) What is the purpose of passing dry water through the air aspirator? (1 mark)
- (ii) State and explain the observation that would be made in tube Y after sometimes. (1 mark)
- (iii) The sample of nitrogen collected at point Y had greater density than expected. What conclusion could be made about the gas? (1 mark)
- (iv) The table below gives the rate of decay of a radioactive element G.

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William Briggs



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