

MINISTRY OF EDUCATION
SECONDARY ENGAGEMENT PROGRAMME
GRADE 10
CHEMISTRY

WEEK 9

LESSON 2 - WORKSHEET

State the type of chemical reaction that occurs in each equation shown below.

- a. $\text{Cu}_{(s)} + \text{O}_{2(g)} \rightarrow \text{CuO}_{(s)}$
- b. $\text{H}_2\text{O}_{(l)} \rightarrow \text{H}_{2(g)} + \text{O}_{2(g)}$
- c. $\text{Fe}_{(s)} + \text{H}_2\text{O}_{(l)} \rightarrow \text{H}_{2(g)} + \text{Fe}_3\text{O}_4_{(s)}$
- d. $\text{AsCl}_{3(aq)} + \text{H}_2\text{S}_{(aq)} \rightarrow \text{As}_2\text{S}_{3(s)} + \text{HCl}_{(aq)}$
- e. $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}_{(s)} \rightarrow \text{CuSO}_{4(aq)} + \text{H}_2\text{O}_{(l)}$
- f. $\text{Fe}_3\text{O}_4_{(s)} + \text{H}_{2(g)} \rightarrow \text{Fe}_{(s)} + \text{H}_2\text{O}_{(l)}$
- g. $\text{CaCO}_{3(s)} \rightarrow \text{CaO}_{(s)} + \text{CO}_{2(g)}$
- h. $\text{Fe}_{(s)} + \text{S}_{(s)} \rightarrow \text{FeS}_{(s)}$
- i. $\text{H}_2\text{S}_{(aq)} + \text{KOH}_{(aq)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{K}_2\text{S}_{(aq)}$
- j. $\text{NaCl}_{(s)} \rightarrow \text{Na}_{(s)} + \text{Cl}_{(s)}$
- k. $\text{Al}_{(s)} + \text{H}_2\text{SO}_4_{(aq)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{Al}_2(\text{SO}_4)_3_{(aq)}$
- l. $\text{H}_3\text{PO}_4_{(aq)} + \text{NH}_4\text{OH}_{(aq)} \rightarrow \text{H}_2\text{O}_{(l)} + (\text{NH}_4)_2\text{PO}_4_{(aq)}$
- m. $\text{C}_3\text{H}_8_{(g)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$
- n. $\text{Al}_{(s)} + \text{O}_{2(g)} \rightarrow \text{Al}_2\text{O}_3_{(s)}$
- o. $\text{CH}_4_{(g)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$
- p. $\text{K}_2\text{SO}_4_{(aq)} + \text{BaCl}_{2(aq)} \rightarrow \text{KCl}_{(aq)} + \text{BaSO}_4_{(s)}$
- q. $\text{C}_3\text{H}_8_{(g)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$
- r. $\text{Ca}(\text{OH})_2_{(aq)} + \text{NH}_4\text{Cl}_{(aq)} \rightarrow \text{NH}_4\text{OH}_{(aq)} + \text{CaCl}_{2(aq)}$
- s. $\text{V}_2\text{O}_3_{(s)} + \text{Ca}_{(s)} \rightarrow \text{CaO}_{(s)} + \text{V}_{(s)}$
- t. $\text{Na}_{(s)} + \text{ZnCl}_{2(aq)} \rightarrow \text{NaCl}_{(aq)} + \text{Zn}_{(s)}$
- u. $\text{C}_3\text{H}_8\text{O}_{2(s)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$
- v. $\text{Ca}_{(s)} + \text{N}_{2(g)} \rightarrow \text{Ca}_3\text{N}_{2(s)}$
- w. $\text{Fe}_3\text{O}_4_{(s)} + \text{H}_{2(g)} \rightarrow \text{Fe}_{(s)} + \text{H}_2\text{O}_{(l)}$
- x. $\text{C}_{12}\text{H}_{26(s)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$
- y. $\text{BN}_{(s)} + \text{F}_{2(g)} \rightarrow \text{BF}_{3(s)} + \text{N}_{2(g)}$
- z. $\text{C}_{12}\text{H}_{26(s)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$

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Table of Contents Answers 100 Reproducible Activities Chemistry

1. Understanding the eBook Answers 100 Reproducible Activities Chemistry
 - The Rise of Digital Reading Answers 100 Reproducible Activities Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers 100 Reproducible Activities Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers 100 Reproducible Activities Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers 100 Reproducible Activities Chemistry
 - Personalized Recommendations
 - Answers 100 Reproducible Activities Chemistry User Reviews and Ratings
 - Answers 100 Reproducible Activities Chemistry and Bestseller Lists
5. Accessing Answers 100 Reproducible Activities Chemistry Free and Paid eBooks
 - Answers 100 Reproducible Activities Chemistry Public Domain eBooks
 - Answers 100 Reproducible Activities Chemistry eBook Subscription Services
 - Answers 100 Reproducible Activities Chemistry Budget-Friendly Options

6. Navigating Answers 100 Reproducible Activities Chemistry eBook Formats
 - ePUB, PDF, MOBI, and More
 - Answers 100 Reproducible Activities Chemistry Compatibility with Devices
 - Answers 100 Reproducible Activities Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers 100 Reproducible Activities Chemistry
 - Highlighting and Note-Taking Answers 100 Reproducible Activities Chemistry
 - Interactive Elements Answers 100 Reproducible Activities Chemistry
8. Staying Engaged with Answers 100 Reproducible Activities Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers 100 Reproducible Activities Chemistry
9. Balancing eBooks and Physical Books Answers 100 Reproducible Activities Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers 100 Reproducible Activities Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers 100 Reproducible Activities Chemistry
 - Setting Reading Goals Answers 100 Reproducible Activities Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers 100 Reproducible Activities Chemistry
 - Fact-Checking eBook Content of Answers 100 Reproducible Activities Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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