

MINISTRY OF EDUCATION
SECONDARY ENGAGEMENT PROGRAMME
GRADE 10
CHEMISTRY

WEEK 9

LESSON 2 - WORKSHEET

State the type of chemical reaction that occurs in each equation shown below.

- a. $\text{___ Cu}_{(s)} + \text{___ O}_{2(g)} \rightarrow \text{___ CuO}_{(s)}$
- b. $\text{___ H}_2\text{O}_{(g)} \rightarrow \text{___ H}_2(g) + \text{___ O}_{2(g)}$
- c. $\text{___ Fe}_{(s)} + \text{___ H}_2\text{O}_{(g)} \rightarrow \text{___ H}_2(g) + \text{___ Fe}_3\text{O}_4(s)$
- d. $\text{___ AsCl}_3(aq) + \text{___ H}_2\text{S}_{(aq)} \rightarrow \text{___ As}_2\text{S}_3(s) + \text{___ HCl}_{(aq)}$
- e. $\text{___ CuSO}_4 \cdot 5 \text{H}_2\text{O}_{(s)} \rightarrow \text{___ CuSO}_4(s) + \text{___ H}_2\text{O}_{(g)}$
- f. $\text{___ Fe}_2\text{O}_3(s) + \text{___ H}_2(g) \rightarrow \text{___ Fe}_{(s)} + \text{___ H}_2\text{O}_{(g)}$
- g. $\text{___ CaCO}_3(s) \rightarrow \text{___ CaO}_{(s)} + \text{___ CO}_{2(g)}$
- h. $\text{___ Fe}_{(s)} + \text{___ S}_{8(s)} \rightarrow \text{___ FeS}_{(s)}$
- i. $\text{___ H}_2\text{S}_{(aq)} + \text{___ KOH}_{(aq)} \rightarrow \text{___ H}_2\text{O}_{(l)} + \text{___ K}_2\text{S}_{(aq)}$
- j. $\text{___ NaCl}_{(s)} \rightarrow \text{___ Na}_{(s)} + \text{___ Cl}_{2(g)}$
- k. $\text{___ Al}_{(s)} + \text{___ H}_2\text{SO}_4(aq) \rightarrow \text{___ H}_2(g) + \text{___ Al}_2(\text{SO}_4)_3(aq)$
- l. $\text{___ H}_3\text{PO}_4(aq) + \text{___ NH}_4\text{OH}_{(aq)} \rightarrow \text{___ H}_2\text{O}_{(l)} + \text{___ (NH}_4)_3\text{PO}_4(aq)$
- m. $\text{___ C}_3\text{H}_8(g) + \text{___ O}_{2(g)} \rightarrow \text{___ CO}_{2(g)} + \text{___ H}_2\text{O}_{(g)}$
- n. $\text{___ Al}_{(s)} + \text{___ O}_{2(g)} \rightarrow \text{___ Al}_2\text{O}_3(s)$
- o. $\text{___ CH}_4(g) + \text{___ O}_{2(g)} \rightarrow \text{___ CO}_{2(g)} + \text{___ H}_2\text{O}_{(g)}$
- p. $\text{___ K}_2\text{SO}_4(aq) + \text{___ BaCl}_2(aq) \rightarrow \text{___ KCl}_{(aq)} + \text{___ BaSO}_4(s)$
- q. $\text{___ C}_3\text{H}_{12}(g) + \text{___ O}_{2(g)} \rightarrow \text{___ CO}_{2(g)} + \text{___ H}_2\text{O}_{(g)}$
- r. $\text{___ Ca(OH)}_2(aq) + \text{___ NH}_4\text{Cl}_{(aq)} \rightarrow \text{___ NH}_4\text{OH}_{(aq)} + \text{___ CaCl}_2(aq)$
- s. $\text{___ V}_2\text{O}_5(s) + \text{___ Ca}_{(s)} \rightarrow \text{___ CaO}_{(s)} + \text{___ V}_{(s)}$
- t. $\text{___ Na}_{(s)} + \text{___ ZnI}_2(aq) \rightarrow \text{___ NaI}_{(aq)} + \text{___ Zn}_{(s)}$
- u. $\text{___ C}_7\text{H}_6\text{O}_3(s) + \text{___ O}_{2(g)} \rightarrow \text{___ CO}_{2(g)} + \text{___ H}_2\text{O}_{(g)}$
- v. $\text{___ Ca}_{(s)} + \text{___ N}_{2(g)} \rightarrow \text{___ Ca}_3\text{N}_2(s)$
- w. $\text{___ Fe}_2\text{O}_3(s) + \text{___ H}_2(g) \rightarrow \text{___ Fe}_{(s)} + \text{___ H}_2\text{O}_{(g)}$
- x. $\text{___ C}_{15}\text{H}_{30}(g) + \text{___ O}_{2(g)} \rightarrow \text{___ CO}_{2(g)} + \text{___ H}_2\text{O}_{(g)}$
- y. $\text{___ BN}_{(s)} + \text{___ F}_{2(g)} \rightarrow \text{___ BF}_3(g) + \text{___ N}_{2(g)}$
- z. $\text{___ C}_{12}\text{H}_{26}(g) + \text{___ O}_{2(g)} \rightarrow \text{___ CO}_{2(g)} + \text{___ H}_2\text{O}_{(g)}$

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