

Chemistry 1st Semester Final Exam Review – ANSWER KEY

READ THE FOLLOWING PARAGRAPH BEFORE YOU START! This worksheet is meant to be a "practice test" to help you prepare for your final exam.

Suggestions on how to prepare: I would first order my notes and homework by the dates on them. Then I would look at the topic statement below (the capital letter phrases) for the first topic. Find this topic and all related information or materials in your possession and study them thoroughly! I would suggest you re-read your textbook/notes and re-work old problems. When you feel you have the topic well reviewed and understood THEN work the topic problems or questions on this sheet. If you don't get the answers right, go back and spend more time reviewing your old materials. If you do get them right, go on to the next topic and repeat the process. After you do all topics this way, start over!

As always, remember to use and report answers to the correct significant figures and to include units!

FUNDAMENTALS OF CHEMISTRY: MATTER AND CHANGE

- Define chemistry.
- Discuss the steps in the scientific method.
- Which of the following observations is evidence used to imply that a chemical reaction has occurred? **a – f**
 - a. gas evolution
 - b. a precipitate forms (**solid formed from aqueous solutions**)
 - c. the color changes
 - d. the temperature changes
 - e. the odor changes
 - f. a sound is produced (**i.e. fireworks**)
 - g. change in taste (**but lets not taste anything during lab class!**)
- Describe a physical property vs. a chemical property.
- What is the difference between a physical change and a chemical change?
- Identify each type of change (physical or chemical).
 - a. boiling water **P**
 - b. rusting iron **C**
 - c. burning candles **C** (**produces new products → $H_2O + CO_2$**)
 - d. breaking glass **P**
 - e. melting wax **P** (**solid to liquid – change in state only**)
- Be able to classify matter: mixtures (homogeneous, heterogeneous) vs. pure substances (compounds or elements)
- Identify each as an element, compound or mixture.
 - a. lead - **element**
 - b. water - **compound**
 - c. kool-aid - **mixture**
 - d. table salt - **compound**
 - e. soda - **mixture**
 - f. oxygen - **element**
- What are the characteristic of a metal, metalloid and non-metal? Identify each type of element:
 - a. Zn - **metal**
 - b. As - **metalloid**
 - c. S - **nonmetal**
 - d. Ne - **nonmetal**
 - e. Na - **metal**

MEASUREMENTS AND CALCULATIONS (Sig. Figs., Scientific Notation, SI Units)

- What are the rules/guidelines of determining significant figures?
- How many significant figures in: **4 sig. figs. in all (a. – e.)**
 - a. 1234
 - b. 123.0
 - c. 123400
 - d. 0.001234
 - e. 0.01230
- Report the above problems (a-e) in correct scientific notation.
 - o **a. 1.234×10^3**
 - b. 1.230×10^2**
 - d. 1.234×10^{-3}**
- What are the common SI Units? Know the prefixes from milli to kilo and be able to convert between units in the metric system.
- What is the difference between mass and weight?
- Multiply: $5.98 \text{ mm} \times 3.2 \text{ mm} = \mathbf{19 \text{ mm}}$
- Add $3.258 \text{ cm} + 12.0 \text{ cm} = \mathbf{15.3 \text{ cm}}$

Answer To 1st Semester Chemistry Exam Review

Richard H. Langley, John Moore



Answer To 1st Semester Chemistry Exam Review:

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