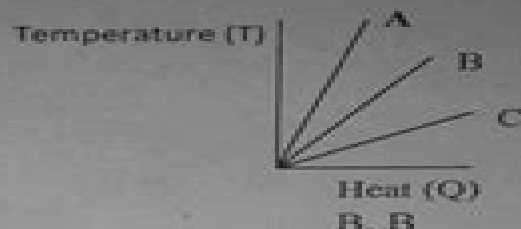


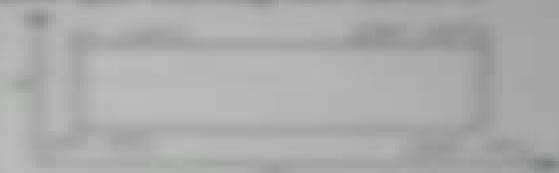
Q. 21 Which of the substances A, B or C has the highest specific heat? The temperature vs heat graph is



A

Heat (Q)
B, B

- Q. 22 The specific heat capacity of a substance is defined as
1. the amount of heat required to raise the temperature of 1 kg of the substance by 1°C
 2. the amount of heat required to raise the temperature of 1 g of the substance by 1°C
 3. the amount of heat required to raise the temperature of 1 kg of the substance by 1 K
 4. the amount of heat required to raise the temperature of 1 g of the substance by 1 K
- Q. 23 Which one is not a thermodynamic process?
1. Isothermal
 2. Adiabatic
 3. Isobaric
 4. Isochoric
- Q. 24 The work done in an isothermal process is
1. $nRT \ln \frac{V_2}{V_1}$
 2. $nRT \ln \frac{P_2}{P_1}$
 3. $nRT \ln \frac{T_2}{T_1}$
 4. $nRT \ln \frac{P_2}{P_1}$



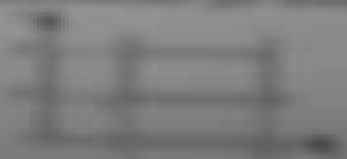
A. 10 J

B. 20 J

C. 30 J

D. 40 J

- Q. 25 Heat energy added to a system under constant volume conditions appears as
1. Work done by the system
 2. Work done on the system
 3. Increase in internal energy
 4. Increase in temperature
- Q. 26 The change in internal energy can be defined as
1. $Q - W$
 2. $W - Q$
 3. $Q + W$
 4. $Q - W$
- Q. 27 A gas expands at 1 atm and 100°C. The work done is
1. 10 J
 2. 100 J
 3. 1000 J
 4. 10000 J
- Q. 28 The work done by an ideal gas during an isothermal process is



A. 10 J

B. 20 J

C. 30 J

D. 40 J

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Thermodynamics of the Steam-engine and Other Heat-engines Cecil Hobart Peabody, 1898 **An Outline of the Theory of Thermodynamics** Edgar Buckingham, 1900 **Computational Thermodynamics of Materials** Zi-Kui Liu, Yi Wang, 2016-06-30 Integrates fundamental concepts with experimental data and practical applications including worked examples and end of chapter problems **Ice and Refrigeration**, 1892 *Industrial Management* John R. Dunlap, 1899

Industrial Management, 1899 **Lightning**, 1891 **The Physical Basis of Biochemistry** Peter R. Bergethon, 2013-03-14 The objective of this book is to provide a unifying approach to the study of biophysical chemistry for the advanced undergraduate who has had a year of physics, organic chemistry, calculus, and biology. This book began as a revised edition of *Biophysical Chemistry: Molecules to Membranes* which Elizabeth Simons and I coauthored. That short volume was written in an attempt to provide a concise text for a one semester course in biophysical chemistry at the graduate level. The experience of teaching biophysical chemistry to biologically oriented students over the last decade has made it clear that the subject requires a more fundamental text that unifies the many threads of modern science: physics, chemistry, biology, mathematics, and statistics. This book represents that effort. This volume is not a treatment of modern biophysical chemistry with its rich history and many controversies; although a book on that topic is also needed. *The Physical Basis of Biochemistry* is an introduction to the philosophy and practice of an interdisciplinary field in which biological systems are explored using the quantitative perspective of the physical scientist. I have three primary objectives in this volume: one to provide a unifying picture of the interdisciplinary threads from which the tapestry of biophysical studies is woven; two to provide an insight into the power of the modeling approach to scientific investigation; and three to communicate

a sense of excitement for the activity and wholesome argument that characterize this field of study Factory and Industrial Management John Robertson Dunlap, Arthur Van Vlassingen, John M. Carmody, 1899 **Thermodynamics of Beta-peptides in Folding, Association, and Adsorption Processes** Clark Austin Miller, 2008 **Industrial Refrigeration**, 1896

Fundamentals of Chemical Engineering Thermodynamics Themis Matsoukas, 2012-10-02 The Clear Well Organized Introduction to Thermodynamics Theory and Calculations for All Chemical Engineering Undergraduate Students This text is designed to make thermodynamics far easier for undergraduate chemical engineering students to learn and to help them perform thermodynamic calculations with confidence Drawing on his award winning courses at Penn State Dr Themis Matsoukas focuses on why as well as how He offers extensive imagery to help students conceptualize the equations illuminating thermodynamics with more than 100 figures as well as 190 examples from within and beyond chemical engineering Part I clearly introduces the laws of thermodynamics with applications to pure fluids Part II extends thermodynamics to mixtures emphasizing phase and chemical equilibrium Throughout Matsoukas focuses on topics that link tightly to other key areas of undergraduate chemical engineering including separations reactions and capstone design More than 300 end of chapter problems range from basic calculations to realistic environmental applications these can be solved with any leading mathematical software Coverage includes Pure fluids PVT behavior and basic calculations of enthalpy and entropy Fundamental relationships and the calculation of properties from equations of state Thermodynamic analysis of chemical processes Phase diagrams of binary and simple ternary systems Thermodynamics of mixtures using equations of state Ideal and nonideal solutions Partial miscibility solubility of gases and solids osmotic processes Reaction equilibrium with applications to single and multiphase reactions *The Second Law of Thermodynamics* Sadi Carnot, Rudolf Clausius, 1899 Reflections on the motive power of heat by Sadi Carnot On the motive power of heat and on the laws which can be deduced from it for the theory of heat by R Clausius The dynamical theory of heat By W Thomson Lord Kelvin

Engineering Magazine, 1899 **The Electrical Journal**, 1896 **Energy Crisis** Naunihal Singh, 2001 *The Electrician*, 1896 *Off-Gas Purification* Annika Graftschafter, Susanne Lux, Matthäus Siebenhofer, 2023-06-06 When doing in the off gas purification business you will pretty soon register that you do not act in an isolated box You have to make yourself familiar with the interplay of your emission problem and the environment and you have to apply a broad view of the subject We can hardly make a forecast on your first steps in this business except that we want you to succeed Therefore we want to offer engineers and graduate students the basic tools for discussing air pollution problems and for deducing strategies for process and equipment design in off gas purification covering the whole span from the basics to dedusting absorption adsorption and redox processes The didactic concept of the work is to attract students with a learning by doing strategy We discuss the problems the solver strategies and the solvers The problem solver proposals address a multitude of pollution control technologies The work is a compact off gas purification guide for practitioners and students by presenting basics as

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