

Name: \_\_\_\_\_ Course & Year: \_\_\_\_\_ Score: \_\_\_\_\_  
 Schedule: \_\_\_\_\_ Date Performed: \_\_\_\_\_ Date Submitted: \_\_\_\_\_  
 Laboratory Instructor: \_\_\_\_\_ Group Number: \_\_\_\_\_

Activity No. 4  
 WORKSHEET

CALORIMETRY AND SPECIFIC HEAT

Data, results, and calculations: Complete the table below, assume that the metal used is iron

(a) Weight of the calorimeter (inner vessel)	9.15 g
(b) Weight of Calorimeter + Water	36.70 g
(c) Weight of water, (b) - (a)	
(d) Initial temperature of water, $(t)_{\text{w}}$	28 °C
(e) Initial temperature of aluminum calorimeter, $(t)_{\text{cal}}$	
(f) Initial temperature of metal $(t)_{\text{metal}}$	98.0 °C
(g) Weight of metal sample, $m_{\text{metal}}$	7.5 g
(h) Final temperature of water, $(t)_{\text{w}}$	32.5 °C
(i) Final temperature of calorimeter, $(t)_{\text{cal}}$	
(j) Final temperature of metal	
(k) Heat absorbed by the calorimeter, $q_{\text{cal}}$	
(l) Heat absorbed by the water, $q_{\text{w}}$	
(m) Experimental specific heat of metal sample	
(n) Theoretical specific heat of metal sample	0.460 cal/g°C
(o) Percent Error	%

GUIDE QUESTIONS

1	What is the purpose of stirring the contents of the calorimeter before the final temperature is taken?	
2	Why is it required to tap the hot metal with a tissue paper before placing the hot metal in the calorimeter?	
3	Why should the transferring of the hot metal from the beaker to the calorimeter be done quickly?	

# Answer Sheet For Calorimetry Specific Heat

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A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the text 'United States. National Bureau of Standards'.

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## **Table of Contents Answer Sheet For Calorimetry Specific Heat**

1. Understanding the eBook Answer Sheet For Calorimetry Specific Heat
  - The Rise of Digital Reading Answer Sheet For Calorimetry Specific Heat
  - Advantages of eBooks Over Traditional Books
2. Identifying Answer Sheet For Calorimetry Specific Heat
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answer Sheet For Calorimetry Specific Heat
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Sheet For Calorimetry Specific Heat
  - Personalized Recommendations
  - Answer Sheet For Calorimetry Specific Heat User Reviews and Ratings
  - Answer Sheet For Calorimetry Specific Heat and Bestseller Lists
5. Accessing Answer Sheet For Calorimetry Specific Heat Free and Paid eBooks
  - Answer Sheet For Calorimetry Specific Heat Public Domain eBooks
  - Answer Sheet For Calorimetry Specific Heat eBook Subscription Services

- Answer Sheet For Calorimetry Specific Heat Budget-Friendly Options
- 6. Navigating Answer Sheet For Calorimetry Specific Heat eBook Formats
  - ePub, PDF, MOBI, and More
  - Answer Sheet For Calorimetry Specific Heat Compatibility with Devices
  - Answer Sheet For Calorimetry Specific Heat Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answer Sheet For Calorimetry Specific Heat
  - Highlighting and Note-Taking Answer Sheet For Calorimetry Specific Heat
  - Interactive Elements Answer Sheet For Calorimetry Specific Heat
- 8. Staying Engaged with Answer Sheet For Calorimetry Specific Heat
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answer Sheet For Calorimetry Specific Heat
- 9. Balancing eBooks and Physical Books Answer Sheet For Calorimetry Specific Heat
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answer Sheet For Calorimetry Specific Heat
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Sheet For Calorimetry Specific Heat
  - Setting Reading Goals Answer Sheet For Calorimetry Specific Heat
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Sheet For Calorimetry Specific Heat
  - Fact-Checking eBook Content of Answer Sheet For Calorimetry Specific Heat
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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