

TEST YOUR SELF

TEST YOURSELF 2.1 (THEORIES AND EXPERIMENTS RELATED TO STRUCTURE OF ATOMS)

i) Do you know any element having no neutrons in its atoms?

Ans: Yes the isotope of hydrogen (Protium, ${}^1_1\text{H}$) has no neutron. It has one proton in its nucleus and one electron revolving around it.

ii) Who discovered an electron, a proton and a neutron?

Ans: a. In 1886, Goldstein discovered positively charged particles called protons.

b. In 1897, J.J. Thomson discovered the negatively charged particles called electrons.

c. In 1932, James Chadwick discovered the neutral particles called neutrons.

iii) How does electron differ from a neutron?

Ans: Difference between electron and neutron.

Electron	Neutron
i. Electron is the negatively charged particle.	i. Neutron is the neutral particle.
ii. Electron revolves around the nucleus.	ii. Neutron is present in the nucleus.
iii. Mass of electron is 5.486×10^{-4} g or 9.106×10^{-28} g.	iii. Mass of neutron is 1.0087 amu or 1.674×10^{-24} g.

iv) Explain how anode rays are formed from the gas taken in the discharge tube?

Ans: Anode rays do not originate from the anode. In fact these rays are produced when the cathode rays or electrons collide with the residual gas molecules present in the discharge tube and ionize them as follows:



TEST YOURSELF 2.2 RUTHERFORD'S ATOMIC MODEL AND BOHR'S ATOMIC MODEL

i) How was it proved that the whole mass of an atom is located at its centre?

Ans: Rutherford bombarded alpha particles on a 0.00004 cm thick gold foil. Almost all the particles passed through the foil undeflected, only few bounced back. The complete rebound of few alpha particles showed that the nucleus is an extremely small positively charged part. It is situated at the center of an atom and it carries nearly the whole mass of an atom.

ii) How was it shown that atomic nuclei are positively charged?

Ans: Alpha particles are helium nuclei (He^{2+}) i.e. doubly positively charged particles. The deflection and rebounding of a few particles showed that atomic nuclei (centre of an atom) are positively charged.

iii) Name the particles which determine the mass of an atom.

Ans: Protons and neutrons present in the nucleus determine the mass of an atom. The sum of number of protons and neutrons is equal to mass of an atom.

iv) What is the classical theory of radiation? How does it differ from quantum theory?

Ans:

Classical theory of radiation	Quantum theory of radiation
According to classical theory of radiation electrons being the charged particles should release or emit energy continuously and they should ultimately fall into the nucleus.	According to quantum theory of radiation, energy is not emitted or absorbed continuously but it is emitted or absorbed in the form of small energy packets or bundles known as quantum or

Quiz Yourself Chapter 8 Introductcion To Chemistry

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Student's Guide to Introduction to Chemical Principles by Edward I. Peters, 2d Ed Peter P. Berlow, 1978 **Aggregate Base and Surfacing Inspection**, 1986 **Forensic Science** Jay A. Siegel, Kathy Mirakovits, 2006-09-07 Forensic Science

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Educart ICSE Class 10 Question Bank 2025 Biology One Shot for 2024-25 Exam Educart, Sir Tarun Rupani, 2024-06-28

Study Guide for Chemical Principles, Fourth Edition, by Dickerson, Gray, Darensbourg, and Darensbourg Patricia L. Samuel, 1984

Basic Chemistry G. William Daub, William S. Seese, 1996 Written in a style and language that users without science backgrounds can understand This best selling introduction to the basic principles of chemistry draws on the reader's own experiences through analogies and cartoons to learn difficult concepts The clear systematic thinking approach to problem solving has also been highly praised by reviewers and users alike Countdown sections in each chapter consisting of five review questions keyed to previous material provide readers with a basis for material introduced in the new chapter Study exercises found immediately after new topics are introduced reinforce chapter problem material You and Chemistry marginal application icon relates chemistry to the real world End of chapter essays entitled Elements and Compounds relate the applications of specific elements or compounds to the reader's life

Fachenglisch für Laborberufe Steven L. Hanft, 2015-09-15 Mit fortschreitender Globalisierung von Waren und Dienstleistungen holt an immer mehr Arbeitsplätzen in Chemie Pharma und Biotech Branche die englische Sprache Einzug In der Schule hat man zwar gelernt sich bei Alltagsthemen zu unterhalten aber wenn es darum geht dem Kundendienst am Telefon die Fehlfunktion des teuersten Geräts im Labor zu beschreiben kommt doch so mancher ins Schwitzen Nach einer Einführung in der die wichtigsten Besonderheiten der englischen Sprache aus Sicht eines deutschen Sprechers rekapituliert werden behandelt der Autor in 14 Lektionen Schritt für Schritt den Spezialwortschatz und fachspezifische Sprach- und Schreibformen Die Themen reichen von mathematischen Ausdrücken über chemische Nomenklatur Biomoleküle Versuchstiere und Prozesstechnik bis hin zum Umgang mit Regulierungsbehörden und Audits Gesprächeituationen wie der Anruf beim Kundendienst die Vorstellung beim neuen Chef oder das Kundengespräch am Messestand werden analysiert und eingebettet Mit direktem Bezug zur Berufspraxis geht dieser Sprachführer über herkömmliche Englischkurse weit hinaus und bietet wertvolle Hilfe für alle die im Beruf besser Englisch sprechen wollen Auch für den fachbezogenen Sprachunterricht an Fachschulen und Hochschulen ist dieses Buch bestens geeignet Komplett mit Übungen Tests und Rezepten wie man die häufigsten Fehler vermeidet Das Buch ist auch als eBook mit Audiounterstützung erhältlich

Introduction to Chemistry Darrell D.

Ebbing, 1995-08

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Introduction to Biology Alan Axelrod, 1999 A complete guide with questions answers and practice tests in the field of biology

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Living Chemistry David A. Ucko, 1986

How to Prepare for the GED High School Equivalency Exam Murray Rockowitz, Brownstein, Max Peters, Bolton, 2001 Provides full length practice exams with answers and explanations

and includes reviews of all test areas from writing skills to science **How to Prepare for the GED with CD-ROM** Murray Rockowitz, Samuel C. Brownstein, Ira K. Wolf, Max Peters, 2002-01-30 Updated to reflect the recent major changes in the high school equivalency exam this manual presents a full length diagnostic test with answer keys answer analyses and self appraisal charts The diagnostic is followed by reviews of all GED test subjects which emphasize proficiency in high school level math reading comprehension and writing skill as well as knowledge of social studies arts and literature Two full length practice exams modeled after the new GED exam complete this manual They come with all questions answered and explained The CD ROM packaged with this book offers a computerized practice GED with automatic scoring plus additional subject review material *Student-active Science* Ann P. McNeal, 1997 This volume emerged from an NSF sponsored conference on Inquiry Approaches to Science Teaching held at Hampshire College in June 1996 STUDENT ACTIVE SCIENCE emphasizes that experiencing the process of science is central to the learning of science This book is a collection of articles ideas and models for science education reform and is the result of collaboration between instructors frustrated with the traditional approach to teaching You ll find models and ideas that promote critical thinking and hands on science in the classroom as well as commentary from school wide department wide and individual reform efforts Edexcel Modular , 2010-02-22 Collins New GCSE Maths Edexcel Modular Teacher s Pack Foundation 1 contains everything you need to deliver effective lessons in mathematics with confidence for students working at Grades G to C Fully matched to Edexcel s new GCSE Maths Modular specification these teacher resources offer well differentiated lesson plans and additional support *Nature* , 1895

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