

AP CHEMISTRY MULTIPLE CHOICE QUESTIONS

Unit 8 – Introduction to Equilibrium



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



Reversed
+ Doubled

(A) 95²

(B) 95

(C) (95)^{1/2}

(D) $\frac{1}{95}$

(E) $\frac{1}{(95)^2}$



$C_2H_2(g)$ is reduced by $H_2(g)$ in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- I. Changes in the partial pressure of $H_2(g)$
- II. Changes in the particle size of the platinum catalyst
- III. Changes in the temperature of the reaction system

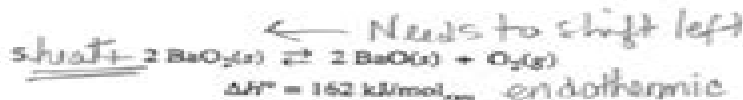
(A) III only

(B) I and II only

(C) I and III only

(D) II and III only

(E) I, II, and III



A sealed rigid vessel contains $BaO_2(s)$ in equilibrium with $BaO(s)$ and $O_2(g)$ as represented by the equation above. Which of the following changes will increase the amount of $BaO_2(s)$ in the vessel?

(A) Removing a small amount of $O_2(g)$ $R \times$

(B) Removing a small amount of $BaO(s)$ X

(C) Adding He gas to the vessel X

(D) Lowering the temperature L

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products?



increase P°
 shift towards fewer
 + make gas



When 4.00 mol each of $X(g)$ and $Y(g)$ are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of $Z(g)$ is produced. What is the value of the equilibrium constant, K_c ?

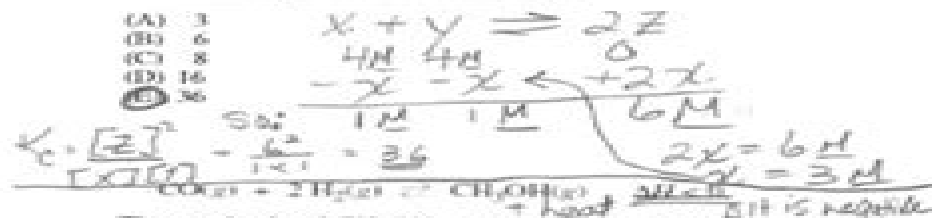
(A) 3

(B) 6

(C) 8

(D) 16

(E) 36



6. The synthesis of $CH_3OH(l)$ from $CO(g)$ and $H_2(g)$ is represented by the equation above. The value of K_c for the reaction at 483 K is 14.5.

increase heat shifts left, decrease K_c
 Which of the following explains the effect on the equilibrium constant, K_c , when the temperature of the reaction system is increased to 650 K?

(A) K_c will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.

(B) K_c will increase because there are more reactant molecules than product molecules.

(C) K_c will decrease because the reaction is exothermic.

(D) K_c is constant and will not change.

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Reino Unido Finlandia Australia etc de primer nivel en las distintas areas escrita integramente en ingles revisa todos los aspectos relacionados con el Sindrome metabolico desde el punto de vista de la fisiopatologia molecular y la epidemiologia genetica obesidad insulinoresistencia dislipidemias hipertension enfermedad cardiovascular y otros nuevos como inflamacion molecular estado protrombotico disfuncion endotelial Todos los capitulos presentan un abstract en ingles y castellano al inicio de cada capitulo con los puntos fundamentales que se van a desarrollar en ese capitulo *Biochemistry and Cell Biology*,2009 *Nordic Pulp & Paper Research Journal* ,2004 Proceedings of the National Academy of Sciences of the United States of America National Academy of Sciences (U.S.),2007 **European Journal of Entomology** ,2008 The Journal of Neuroscience ,2009 **International Directory of Company Histories** Jay P. Pederson,2007-04 This multi volume series provides detailed histories of more than 8 500 of the most influential companies worldwide **The Journal of Cell Biology** ,2007 No 2 pt 2 of November issue each year from v 19 1963 47 1970 and v 55 1972 contain the Abstracts of papers presented at the Annual Meeting of the American Society for Cell Biology 3d 1963 10th 1970 and 12th 1972

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