

# ATOMIC STRUCTURE

| Particle           | Symbol     | Charge                   | Relative charge | Mass                               | Relative mass w.r.t H | e/m ratio |
|--------------------|------------|--------------------------|-----------------|------------------------------------|-----------------------|-----------|
| Electron           | $e^-$      | $-1.6 \times 10^{-19} C$ | -1              | $9.1 \times 10^{-28}$ gram         | $1/1837$              | 1837      |
| Proton             | $p^+$      | $+1.6 \times 10^{-19} C$ | +1              | $1.673 \times 10^{-24}$ g          | 1                     | 1         |
| Neutron            | $n^0$      | Zero                     | Zero            | $1.678 \times 10^{-24}$ g          | 1                     | 0         |
| $\alpha$ -Particle | ${}^2He^4$ | $1.6 \times 10^{-19} C$  | +2              | $4 \times 1.678 \times 10^{-24}$ g | 4                     | $1/2$     |

## # Thomson's atomic model

(Plum pudding / Watermelon model)

- ve charge  $\rightarrow$  uniformly distributed
- ve charge  $\rightarrow$  embedded in cloud of the charge.
- Limitation  $\rightarrow$  can't explain stability

## # Rutherford's experiment

(Discovery of nucleus)

Atom  $\rightarrow$  Nucleus (Proton)  
 $\rightarrow$  Electrons (part ( $e^-$ ))

$$\text{Nucleus} = K(A)^{1/3}$$

$$\frac{(\text{mass})_1}{(\text{mass})_2} = \left(\frac{A_1}{A_2}\right)^{1/3}$$

$$V_{\text{nucleus}} = \frac{4}{3} \pi R^3 \rho$$

$$\frac{V_{\text{nucleus}}}{V_{\text{atom}}} = \frac{A_1}{A_2}$$

$$V_{\text{atom}} \approx 10^{15} \times V_{\text{nucleus}}$$

## # Term associated with element

~~$Z$~~   $\rightarrow$  Mass number ( $P+q$ )  
 $\rightarrow$  Atomic number ( $P$ )

- Isotope  $\rightarrow$   $Z$   $\rightarrow$  same  $\rightarrow$  different
- Isobar  $\rightarrow$   $Z$   $\rightarrow$  different  $\rightarrow$  same
- Isotone  $\rightarrow$   $A-Z = \eta \rightarrow$  same
- Isoelectronic  $\rightarrow$  Same no. of  $e^-$  ( $NO_3^-$ ,  $CO$ )
- Isotopes  $\rightarrow$  Same no. of atom and  $e^-$  ( $CO_2$ ,  $NO_2$ )
- Isotrophies  $\rightarrow$  ( $1-P$ ) or ( $A-2Z$ )  $\rightarrow$  same
- Paramagnetic  $\rightarrow$  Having non-zero unpaired  $e^-$
- Diamagnetic  $\rightarrow$  Having zero unpaired  $e^-$

## # Dual nature of light

### Wave nature

- Wavelength ( $\lambda$ )  
 distance between 2 crest or trough or length of one wave.
- Wave no.  $\bar{v} = \frac{1}{\lambda}$

$$\text{frequency } v = \frac{c}{\lambda}$$

$$c \approx 3 \times 10^8 \text{ m/s}$$

$$\lambda \propto \frac{1}{\bar{v}} \rightarrow \frac{v_1}{v_2} = \frac{\lambda_2}{\lambda_1}$$

### Particle nature

- photon  $\rightarrow$  quanta of light.
- $E = hv$   
 $h \rightarrow$  plank's constant  
 $h = 6.6 \times 10^{-34}$  joule sec

$$E = n \times E_{\text{min}}$$

$\nearrow$   
 integral multiple

$$E = h\bar{v} = \frac{hc}{\lambda}$$

$$\frac{E_1}{E_2} = \frac{v_1}{v_2} \quad \frac{E_1}{E_2} = \frac{\lambda_2}{\lambda_1}$$

## # Photo electric effect

- example of particle nature.

- Surface phenomena

- Single photon release single  $e^-$

$$h\nu = \omega + KE \quad \omega \rightarrow \text{work function}$$

$$h\nu = h\nu_0 + \frac{1}{2}mv_e^2 \quad \nu \rightarrow \text{incident frequency of photon}$$

$$\frac{h\nu}{\lambda} = \frac{hc}{\lambda_0} + \frac{1}{2}mv_e^2 \quad \nu_0 \rightarrow \text{threshold frequency of metal.}$$

$\# \nu > \nu_0 \rightarrow e^- \text{ emitted with KE}$   
 $(\lambda < \lambda_0)$

$\# \nu = \nu_0 \rightarrow e^- \text{ emitted without KE}$   
 $(\lambda = \lambda_0)$

$\# \nu < \nu_0 \rightarrow e^- \text{ not emitted}$   
 $(\lambda > \lambda_0)$

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