

Name Key Class \_\_\_\_\_ Date \_\_\_\_\_

## Atomic Structure

### Subatomic Particles in Neutral Atoms

The table below contains information about several isotopes. Use the information given to fill in the blanks. Assume all atoms are neutral.

| Isotope Name  | Nuclear Symbol           | Atomic Number | Mass Number | # of Protons | # of Neutrons | # of Electrons |
|---------------|--------------------------|---------------|-------------|--------------|---------------|----------------|
| 1. calcium-40 | ${}_{20}^{40}\text{Ca}$  | 20            | 40          | 20           | 20            | 20             |
| 2. iron-56    | ${}_{26}^{56}\text{Fe}$  | 26            | 56          | 26           | 30            | 26             |
| 3. oxygen-18  | ${}_{8}^{18}\text{O}$    | 8             | 18          | 8            | 10            | 8              |
| 4. gold-197   | ${}_{79}^{197}\text{Au}$ | 79            | 197         | 79           | 118           | 79             |

### Subatomic Particles in Ions

The table below contains information about several ions. Use the information given to fill in the blanks.

| Element Name | Ion Symbol       | Atomic Number | Mass Number | # of Protons | # of Neutrons | # of Electrons |
|--------------|------------------|---------------|-------------|--------------|---------------|----------------|
| 5. chlorine  | $\text{Cl}^{-}$  | 17            | 35          | 17           | 18            | 18             |
| 6. silver    | $\text{Ag}^{+}$  | 47            | 107         | 47           | 60            | 46             |
| 7. oxygen    | $\text{O}^{2-}$  | 8             | 16          | 8            | 8             | 10             |
| 8. aluminum  | $\text{Al}^{3+}$ | 13            | 27          | 13           | 14            | 10             |

### Average Atomic Mass

9. Calculate the average atomic mass for neon if its abundance in nature is 90.5% neon-20, 0.3% neon-21, and 9.2% neon-22.

$$\frac{(90.5)(20 \text{ amu}) + (0.3)(21 \text{ amu}) + (9.2)(22 \text{ amu})}{100} = 20.18 \text{ amu}$$

10. Calculate the average atomic mass of silver if 13 out of 25 atoms are silver-107 and 12 out of 25 atoms are silver-109.

$$\frac{(13)(107 \text{ amu}) + (12)(109 \text{ amu})}{25} = 107.96 \text{ amu}$$

# Atomic Structure Practice Answer Key Chemistry

**Chao Zhang**



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