

Atomic Structure



Atomic Number

Element Symbol

Element Name

Atomic Mass

The atomic number represents these two parts of the atom:
number of protons & number of electrons

You can calculate the number of neutrons by subtracting

Atomic Mass minus Atomic Number



Protons:

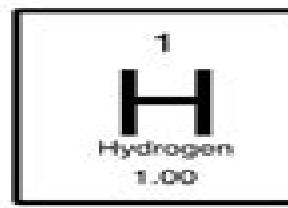
7

Electrons:

7

Neutrons:

7



Protons:

1

Electrons:

1

Neutrons:

0



Protons:

8

Electrons:

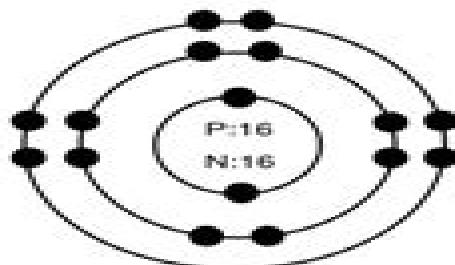
8

Neutrons:

8

BOHR DIAGRAMS

- Draw a circle in the center representing the nucleus.
- Label the number of protons and neutrons.
- Draw the electrons in their energy levels.
 1. The first energy level can hold 2 electrons
 2. The second energy level can hold 8 electrons
 3. The third energy level can hold 8 electrons

Element: SulfurNumber of Valence Electrons: 6

The electrons in the outermost energy level are called:

Valence Electrons

LEWIS DIAGRAMS

A Lewis diagram shows only the symbol with the Valence electrons



Atomic Basics Atomic Structure Worksheet Answers

American Society for Microbiology

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Teaching and Learning Online Franklin S. Allaire, Jennifer E. Killham, 2023-01-10 Teaching and Learning Online Science for Elementary Grade Levels explores the challenges of teaching science virtually. It includes sections on frameworks, teacher journeys and lesson plans aligned with Next Generation Science Standards offering tips, resources and discussion questions for educators and students. *Learning Elementary Chemistry for Class 8 (A.Y. 2023-24) Onward* Dr. R. Goel, 2023-05-20 The series Learning Elementary Chemistry for Classes 6 to 8 has been revised strictly according to the latest curriculum. The content of this series has been developed to fulfill the requirement of all the six domains: Concepts, Processes, Applications, Attitudes, Creativity, and World view of Science to make teaching and learning of Chemistry interesting, understandable, and enjoyable for young minds. This series builds a solid foundation for young learners to prepare them for higher classes. The main strength of the series lies in the subject matter and the experience that a learner will get in solving difficult and complex problems of Chemistry. Emphasis has been laid upon mastering the fundamental principles of Chemistry rather than specific procedures. Unique features of this series are:

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