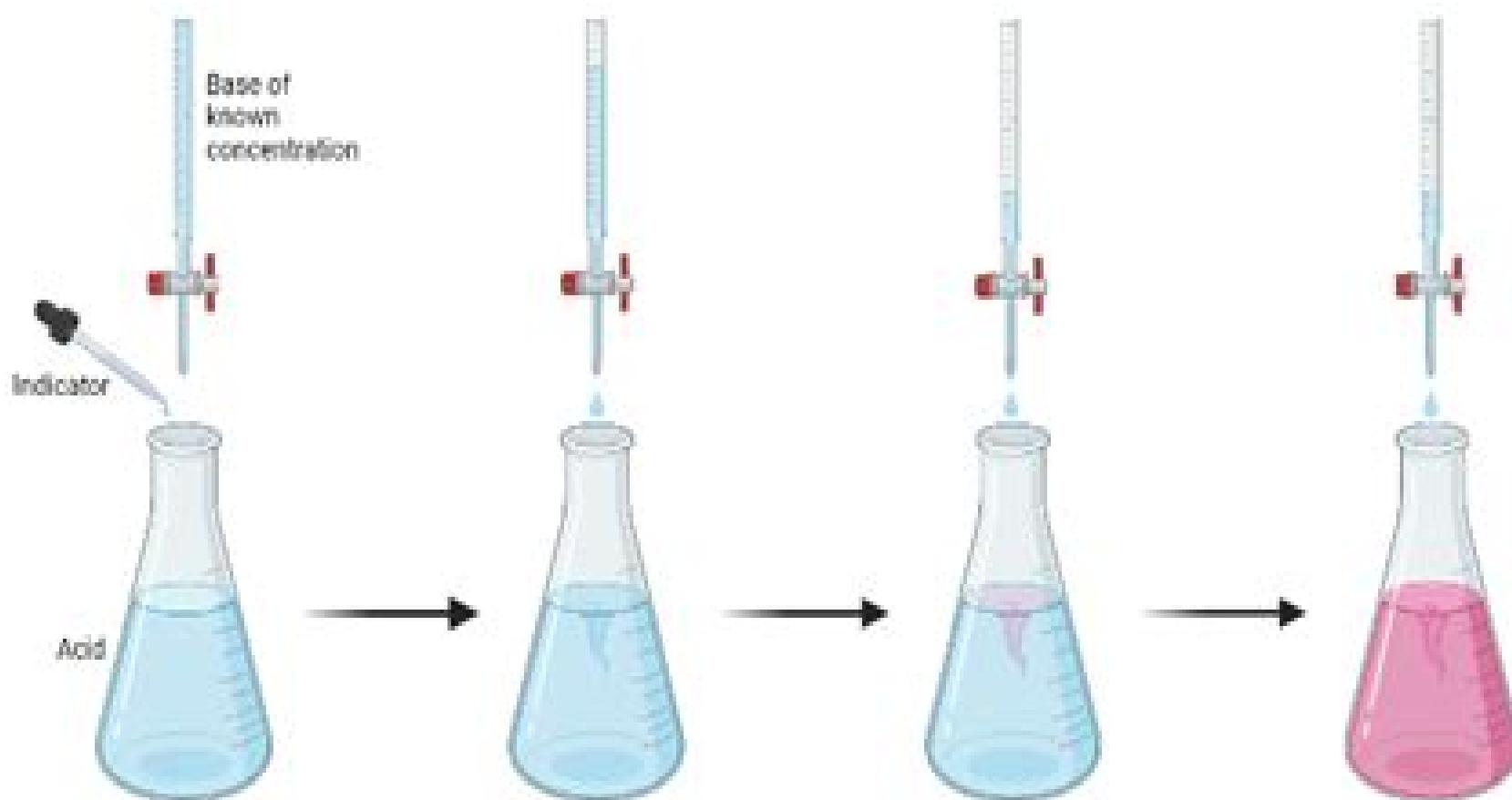


Acid-Base Titration



① A pH indicator is added to the flask containing an aqueous solution of the acid of unknown concentration.

② The base is added slowly (drop by drop) to the acid and mixed repeatedly.

③ Equal amount of OH^- (base) and H^+ (acid) causes indicator to change its color. At this stage we should stop titration.

④ If the amount of OH^- surpasses the amount of H^+ , the solution changes color completely. This should be avoided.

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