

# ATOMIC STRUCTURE

## RUTHERFORD'S MODEL

- Discovered that an atom is mostly made up of empty space
- Center of the atom is a small dense positively charged nucleus
- This nucleus holds the negatively charged particles in the atom

## BOHR'S MODEL

- There are 2 main pieces of an atom:
- Nucleus
- Electrons

## THE OLD MODEL OF THE ATOM



## WAVE-MECHANICAL MODEL

- Electrons are moving so fast that they look like a cloud. Orbitals are regions of most likely electron location



## WHAT IS TRUE TODAY...

- The nucleus is composed of 2 things:
  - PROTONS - positively charged particles
  - NEUTRONS - neutral charged particles
- Electrons are in a cloud - they are in orbitals
  - Electrons DO move around the nucleus in different patterns



| Subatomic Particle | Location in the Atom                     | Charge              | Weight                 |
|--------------------|------------------------------------------|---------------------|------------------------|
| proton             | center (in the nucleus)                  | positive            | 1 amu                  |
| neutron            | center (in the nucleus)                  | neutral (no charge) | 1 amu                  |
| electron           | around the nucleus in clouds or orbitals | negative            | Almost equal to 1/1836 |

### IONS

- If an atom gains or loses an electron, it is called an ion
- This gives the atom a charge:
  - Lose an electron, the atom becomes positive
  - Gain an electron, the atom becomes negative

### ISOTOPES

- An isotope is when an atom is physically different
- An isotope is still the same element, it is just a little different from every other atom of the same element
  - slightly different mass

# Atomic Structure Study Guide

**T. W. Graham Solomons**



## Atomic Structure Study Guide:

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