

# ATOMIC STRUCTURE

## RUTHERFORD'S MODEL

- Discovered that an atom is mostly made up of empty space
- Center of the atom is a small dense positively charged nucleus
- This nucleus holds the negatively charged particles in the atom

## BOHR'S MODEL

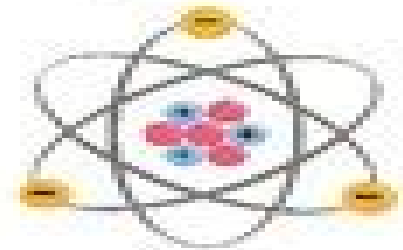
- There are 2 main pieces of an atom:
- Nucleus
- Electrons

### THE OLD MODEL OF THE ATOM



## WAVE-MECHANICAL MODEL

- Electrons are moving so fast that they look like a cloud. Orbitals are regions of most likely electron location



## WHAT IS TRUE TODAY...

- The nucleus is composed of 2 things:
  - PROTONS - positively charged particles
  - NEUTRONS - neutral charged particles
- Electrons are in a cloud - they are in orbitals
  - Electrons DO move around the nucleus in different patterns



Subatomic Particle	Location in the Atom	Charge	Weight
proton	center (in the nucleus)	positive	1 amu
neutron	center (in the nucleus)	neutral (no charge)	1 amu
electron	around the nucleus in clouds or orbitals	negative	Almost equal to neutrons (1/2)

### IONS

- If an atom gains or loses an electron, it is called an ion
- This gives the atom a charge:
  - Lose an electron, the atom becomes positive
  - Gain an electron, the atom becomes negative

### ISOTOPES

- An isotope is when an atom is physically the same, but different
- An isotope is still the same element, it is just a little different from every other atom of the same element
  - slightly different mass

# Atomic Structure Anticipation Guide

**G Orfield**



## Atomic Structure Anticipation Guide:

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Resources in Education, 1984 Serves as an index to Eric reports microform

**Study Guide, Student Edition** Merrill, 1994-07

*Study Guide with Answers to Selected Problems* H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

**Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl** Paul B. Kelter, 1993

**Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry** Katherine J. Denniston, Denniston Katherine, 2006

*Chemistry, Study Guide* James E. Brady, John R. Holum, 1993-02-11 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples

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**Essential Concepts of Chemistry Study Guide** James R. Braun, Sherman, 1999

**Basic Concepts of Chemistry, Study Guide** Leo J. Malone, 2000 The text s three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout

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