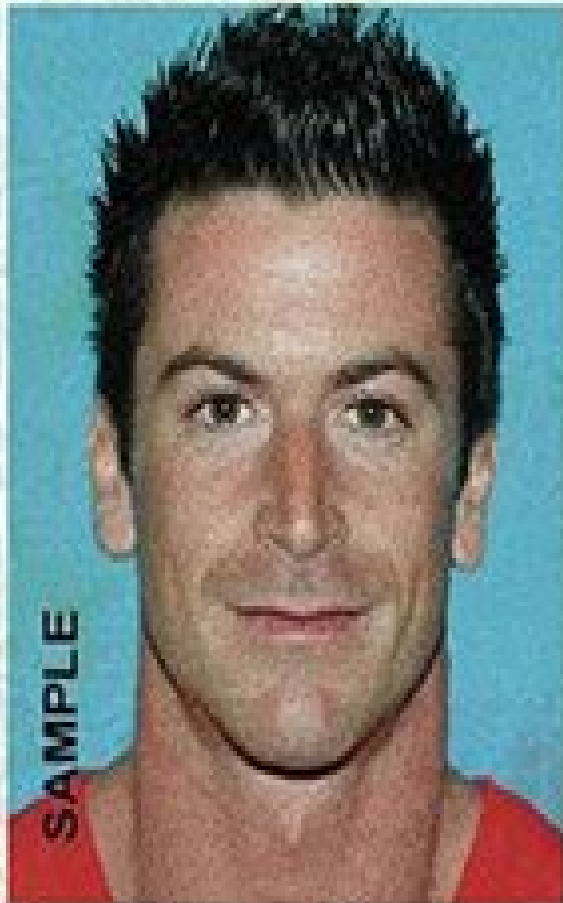


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9a **END** NONE
12 **REST** NONE

4d **DLN** D12345678

3 **DOB** 02/01/1957

1 **SAMPLE**

2 **JELANI**

8 **123 MAIN STREET**
PHOENIX, AZ 85007

4b **EXP** 02/01/2018

4a **ISS** 01/10/2013

15 **SEX** M

18 **EYES** BRO

16 **HGT** 5'-08"

19 **HAIR** BRO

17 **WGT** 185 lb

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02/01/57

Jelani Sample

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5 **DD** 1234567890123456



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web where the salt is the conjugate base of the acid the four parts of the titration curve are described below and you should look to the appropriate text section to see how they are treated pure acid 0 ml of base is added section 17.3.2.1

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of HCl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then $ph = -\log 0.100 = 1.000$ $ph = -\log 0.100 = 1.000$ b titrant volume 12.50 ml

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Volume of NaOH added (mL)	pH
4.6	4.4
4.6	4.5
4.6	4.6
4.6	4.7
4.6	4.8
4.6	4.9
4.6	5.0
4.6	5.1
4.6	5.2
4.6	5.3
4.6	5.4
4.6	5.5
4.6	5.6
4.6	5.7
4.6	5.8
4.6	5.9
4.6	6.0
4.6	6.1
4.6	6.2
4.6	6.3
4.6	6.4
4.6	6.5
4.6	6.6
4.6	6.7
4.6	6.8
4.6	6.9
4.6	7.0
4.6	7.1
4.6	7.2
4.6	7.3
4.6	7.4
4.6	7.5
4.6	7.6
4.6	7.7
4.6	7.8
4.6	7.9
4.6	8.0
4.6	8.1
4.6	8.2
4.6	8.3
4.6	8.4
4.6	8.5
4.6	8.6
4.6	8.7
4.6	8.8
4.6	8.9
4.6	9.0
4.6	9.1
4.6	9.2
4.6	9.3
4.6	9.4
4.6	9.5
4.6	9.6
4.6	9.7
4.6	9.8
4.6	9.9
4.6	10.0

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