

## HONORS CHEM 1st SEMESTER EXAM STUDY GUIDE FALL 2010

**STRATEGY:** Start by reading through your notes to refresh your memory on these topics. Then, use this review sheet as a starting point to identify the areas on which you need to spend more study time. For those areas, go back to homework assignments, quizzes, and reviews to practice more problems. I would also recommend going through all of your tests since these questions are only samples and do not include specific examples of how vocabulary and other conceptual information might appear in a multiple-choice or other format. Remember you can access notes and reviews on my website.

**IMPORTANT INFORMATION:**

- This must be handed in BEFORE taking your final exam.
- If entirely complete, 5 points will be added to your exam grade
- All work must be shown in DETAIL on lined paper

**Introduction to Chemistry – Chapter 1**

1. Identify safe/unsafe lab practices.
2. List pieces and uses of lab equipment.
3. List the requirements for tables and graphs.
4. What is plotted on the x-axis and on the y-axis?
5. Make a density graph.
6. What is water displacement?

**Matter and Change – Chapter 2**

7. Classify the following substances as *solid*, *liquid*, *gas*, or *plasma* based on their properties.
  - a. flexible volume, high KE, particles can disperse freely.
  - b. flexible volume, very high KE, particles are charged.
  - c. fixed volume, very low KE, orderly particles.
  - d. fixed volume, low KE, particles can move past each other.
8. Classify the following as element, compound, heterogeneous mixture, or solution.
  - a. graphite (carbon)
  - b. grape juice
  - c. table salt (NaCl)
  - d. pepper
9. Classify the following as *chemical* or *physical changes*.
  - a. cutting wire
  - b. ripening tomato
  - c. apple slices turning brown
  - d. compressing a gas
10. Classify the following properties as *physical* or *chemical*.
  - a. melts at 68.0°C
  - b. corrosive
  - c. reacts violently with water
  - d. decomposes in air
  - e. magnetic
11. Classify the following properties as *intensive* or *extensive*.
  - a. melting point
  - b. volume
  - c. mass
  - d. reactivity
  - e. density

**VOCAB:** kinetic molecular theory

# 1st Semester Chemistry Study Guide

**M Carnoy**



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*Resources in Education*, 1984      **The Technique of Study in First Year Chemistry** Edward Wallace Johnson, 1930

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## **Table of Contents 1st Semester Chemistry Study Guide**

1. Understanding the eBook 1st Semester Chemistry Study Guide
  - The Rise of Digital Reading 1st Semester Chemistry Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying 1st Semester Chemistry Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 1st Semester Chemistry Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from 1st Semester Chemistry Study Guide
  - Personalized Recommendations
  - 1st Semester Chemistry Study Guide User Reviews and Ratings
  - 1st Semester Chemistry Study Guide and Bestseller Lists
5. Accessing 1st Semester Chemistry Study Guide Free and Paid eBooks
  - 1st Semester Chemistry Study Guide Public Domain eBooks
  - 1st Semester Chemistry Study Guide eBook Subscription Services
  - 1st Semester Chemistry Study Guide Budget-Friendly Options
6. Navigating 1st Semester Chemistry Study Guide eBook Formats

- ePub, PDF, MOBI, and More
- 1st Semester Chemistry Study Guide Compatibility with Devices
- 1st Semester Chemistry Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 1st Semester Chemistry Study Guide
  - Highlighting and Note-Taking 1st Semester Chemistry Study Guide
  - Interactive Elements 1st Semester Chemistry Study Guide
- 8. Staying Engaged with 1st Semester Chemistry Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 1st Semester Chemistry Study Guide
- 9. Balancing eBooks and Physical Books 1st Semester Chemistry Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 1st Semester Chemistry Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 1st Semester Chemistry Study Guide
  - Setting Reading Goals 1st Semester Chemistry Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 1st Semester Chemistry Study Guide
  - Fact-Checking eBook Content of 1st Semester Chemistry Study Guide
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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