

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HBr~~ **hydrobromic acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO₃~~ **nitric acid**
- ~~H₂SO₄~~ **sulfuric acid**
- ~~Strontium hydroxide~~ **LiOH**
- ~~potassium hydroxide~~ **KOH**
- ~~Sodium hydroxide~~ **NaOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~ammonium hydroxide~~ **LiOH**

HClO₄ is PERMANENTLY acidic & never will completely dissociate in water.

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

~~HF~~

~~H₂SO₄~~

~~Mg(OH)₂~~

NaOH

NaHCO₃

~~NaOH~~

Na₂CO₃

3. What is the formula for hydronium? **H₃O⁺**

4. BaCu_2O_6 will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{BaCu}_2\text{O}_6 + \text{H}_2\text{O} \rightarrow \text{BaCu}_2\text{O}_5 + \text{H}^+$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because BaCu_2O_6 is a weak acid.**
- Would you define BaCu_2O_6 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is BaCu_2O_6 acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it's a proton donor.**
- Identify the conjugate base of BaCu_2O_6 : **BaCu_2O_5**
- Identify BaCu_2O_6 as polyprotic, amphoteric, or neither: **polyprotic, because it contains multiple H's.**

5. F^- will react with water and gain a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{F}^- + \text{H}_2\text{O} \rightarrow \text{HF} + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**
- Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**
- In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it's a proton acceptor.**
- Identify the conjugate acid of F^- : **HF**
- Identify F^- as polyprotic, amphoteric, or neither: **neither**

6. HPO_4^{2-} will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{HPO}_4^{2-} + \text{H}_2\text{O} \rightarrow \text{H}_2\text{PO}_4^- + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HPO_4^{2-} is a weak acid.**
- Would you define HPO_4^{2-} as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is HPO_4^{2-} acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it's a proton donor.**
- Identify the conjugate base of HPO_4^{2-} : **PO_4^{3-}**
- Identify HPO_4^{2-} as polyprotic, amphoteric, or neither: **amphoteric, because of both its negative charge (only one H+ to lose) and its H's (can lose or gain an H+).**

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Kenneth W. Whitten, Raymond E. Davis

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