

18.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the K_a and K_b of acids and bases and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization | 5. K_a | 6. K_b |
| 7. ionization constant | 8. pH | |

Key Equation

- | | |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

Part A Completion

Write the equation and/or the numerical value(s) of the K_a and K_b for the acids and bases listed in the table. Also indicate which ionization constant is the stronger one.

1. HNO_3 is a strong acid. K_a is very large.
2. HCl is a strong acid. K_a is very large.
3. H_2SO_4 is a strong acid. K_a is very large.
4. H_2SO_3 is a weak acid. K_a is small.
5. H_2CO_3 is a weak acid. K_a is small.
6. H_2O is a weak acid. K_a is small.
7. H_2O is a weak base. K_b is small.
8. H_2O is a weak acid. K_a is small.
9. H_2O is a weak base. K_b is small.
10. H_2O is a weak acid. K_a is small.
11. H_2O is a weak base. K_b is small.
12. H_2O is a weak acid. K_a is small.
13. H_2O is a weak base. K_b is small.
14. H_2O is a weak acid. K_a is small.
15. H_2O is a weak base. K_b is small.

| | |
|-----|-----------|
| 1. | HNO_3 |
| 2. | HCl |
| 3. | H_2SO_4 |
| 4. | H_2SO_3 |
| 5. | H_2CO_3 |
| 6. | H_2O |
| 7. | H_2O |
| 8. | H_2O |
| 9. | H_2O |
| 10. | H_2O |
| 11. | H_2O |
| 12. | H_2O |
| 13. | H_2O |
| 14. | H_2O |
| 15. | H_2O |

19 2 Hydrogen Ions And Acidity Worksheet Answers

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Company, William Anthony Taylor**



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